

Lab 1- What is around us?

2/1/2024

Focus Questions

- 1) What types of compounds do you know? Give a few examples.
 - Two examples of compounds are ionic and covalent bonds.
- 2) When naming different types of compounds do you need to follow different rules? Explain using examples.
 - For naming covalent bonds first write the names of both elements. Then change the ending of the second one to ide and add the necessary prefixes.
 - $\text{SiCl}_4 \rightarrow$ Silicon Tetrachloride
 - For ionic compounds, write the metal's name first and then the non-metal. Change non-metals ending to ide. Add roman numerals if necessary.
 - $\text{MgBr}_2 \rightarrow$ Magnesium Bromide
 - $\text{CuBr} \rightarrow$ Copper (I) Bromide
- 3) Does it matter in which order you combine your elements when writing chemical formulas for ionic compounds?
 - Yes it does. The cation is typically written first which is the atom which is losing an electron. This means the positive ion is written first and then the negative ion is. The polyatomic ion is always written last.
- 4) Which laws govern the formation of compounds?
 - The law of mass action.

Post-Lab Assessment Questions (20 points):

1. (6 pts) For the following pairs of ions, write the formula of the compound that you would expect them to form:

- a. barium and cyanide $\text{Ba}(\text{CN})_2$
- b. cobalt(III) and sulfur Co_2S_3
- c. iron(III) and sulfate ~~FeSO₄~~ $\text{Fe}_2(\text{SO}_4)_3$

2. (6 pts) Platinum is a transition metal and forms two different ions, Pt^{2+} and Pt^{4+} . Write the formulas for the compounds for each platinum ion with:

- a. chloride ions PtCl_2 , PtCl_4
- b. carbonate ions PtCO_3 , $\text{Pt}(\text{CO}_3)_2$
- c. nitrate ions $\text{Pt}(\text{NO}_3)_2$, $\text{Pt}(\text{NO}_3)_4$

3. (4 pts) Name the following compounds:

- a. N_2O_5 Dinitrogen Pentaoxide
- b. AlCl_3 ~~Aluminum Chloride~~ Aluminium Chloride
- c. K_3PO_4 Tripotassium Phosphate
- d. Mn_2O_7 Manganese (VII) Oxide

4. (4 pts) Write the formulas for the following compounds:

- a. Sulfur tetrachloride SCl_4
- b. Bromine dioxide BrO_2
- c. Nickel(II) nitrate $\text{Ni}(\text{NO}_3)_2$
- d. Calcium sulfide CaS

Hunter Lab Manual Chem 106 Lab

Michael Brown



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