#### CHEMISTRY I EQUATION SHEET

	- 5	olubility Rules		
Negative Ion P	dus Po	sitive ion	Form a Compound which is	
		la ions Soluble		
		nium ion Soluble		
Nitrate + an	y positiv	ve ion Soluble		
Acetate		Ag or Hg or Cu	Soluble	
		Ag Pb Hg or Cu	Not Se	Autoles
bromide or iodide + An	y other	positive ion Soluble		
Sulfate		Ca <sup>3*</sup> ,Sr <sup>3*</sup> ,Ba <sup>3*</sup> , Ra <sup>3*</sup> ,Ag*or Pb <sup>3*</sup>	Not	Solubi
	+ Any other positive ion Soluble			
Sulfide		group 1a ions or ammonium Soluble		
		Be <sup>3*</sup> ,Mg <sup>3*</sup> ,Ca <sup>3*</sup> , Sr <sup>3*</sup> ,Ba <sup>3*</sup> or Ra <sup>3*</sup>	Soluble	
	+ A	ny other positive ion Not Sc	eliulbiler eliulbiler	
Hydroxide	+	group 1a ions or ammonium Soluble		
	+ /4/	ny other positive ion Not Sc	oliubile	
Carbonate, Chromate,		group 1a ions or ammonium Soluble		
Phosphate, or Sulfite + An	y other	positive ion Not Soluble		

Equations			
D = m/v  Molarity = mol V  P, V, = P, V, T, T; Q = mc \Delta T  nRTPV - F = I (1/2) " n = Total Time Half Life Time	c = specific heat d = density I = initial amount F = final amount m = mass M = molarity n = number of moles & number of half lives P = pressure R = ideal gas constant Q = heat energy T = temperature V = volume X = amount of matter left		

	Constants		
	1 otm = 10	03.7 kPa Pressur	
	1 atm = 760 mm Hg		
	R = 8.314 L x kPa / mol*C Uni		
Constant or 0.0	821 L x atm / moi*	C	
Avogadro's number	6.02	×	
Specific Heats (J/g*C)	Water = 4,180 Aluminum = 0.895		
	Copper = 0.387		
	Iron = 0.448		

Activity Series in	
Decreasing	Į.
Activity	
Metals	
Lithium	
Rubidium	
Potassium	
Calcium	
Sodium	
Magnesium	
Aluminum	_
Manganese	
Zime	_
Iron	_
Nickel	
Tim	_
Lead	_
Hydrogen	
Copper	_
Silver	-
Diatinum	
Gold	-
Minmetals	-
PRogramme class	-
Chlorine	-
Bromine	-
iodine	-

HEEI	
Polyatomic	lons
Name Formula	
Ammonium	NH. "1
Acetate	C <sub>2</sub> H <sub>2</sub> O <sub>2</sub> <sup>-1</sup>
Carbonate	CO3 -3
Bicarbonate or Hydrogen carbonate	HCO <sub>3</sub> <sup>-1</sup>
Hypochiorite Chioride.	CIO 1
Chilorite	CIO <sub>2</sub> <sup>-1</sup>
Chlorate	CIO <sub>3</sub> -1
Perchlorate	CIO <sub>4</sub> -1
Hydroxide	OH -1
Cyanide	CIN "
Nitrate	NO <sub>3</sub> ·1
Nitrite	NO <sub>2</sub> ·1
Chromate	CrO, <sup>2</sup>
Oxalate	C2O4-2
Sulfate	SO. 3
Sulfite	SO <sub>3</sub> -2
Hydrogen sulfate	HSO. "
Phosphate	PO <sub>4</sub> -3
Phosphite	PO <sub>3</sub> -9
Hydrogen phosphate	HPO, 2

Metric Conversions			
Umit	Abbr.	Meaning	
Mega-	M	10 *	
killo-	- k	10	
hecto-	h	10 2	
deca-	dk	10 1	
Liter	L	1	
gram	9	1	
meter	m	1	
deci-	d	10-1	
coenti-	C	10 -2	
milli-	m	103	
milicro-	100	10 *	
nano-	n	10 **	

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