

Hewlett-Packard Financial Performance Assessment

Exercise 10.17 has three similar following problems. Use all appropriate and proper scientific units. Use scientific notation for answers (e.g., 1.23×10^4).

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_{18}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$.
 Element Data: $\text{C}_{10}\text{H}_{18}\text{O} = 1.23 \text{ kJ/mol}$ $\text{H}_2\text{O} = -285.8 \text{ kJ/mol}$
 and: $\text{CO}_2 = -393.5 \text{ kJ/mol}$ $\text{C}_{10}\text{H}_{18}\text{O} = 1.23 \text{ kJ/mol}$ or $1.23 \times 10^3 \text{ J/mol}$

a. Answer: $\text{C}_{10}\text{H}_{18}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$ $\Delta H^\circ = -1000 \text{ kJ}$
 or $\Delta H^\circ = -1.00 \times 10^3 \text{ kJ}$

b. Add: $\text{C}_{10}\text{H}_{18}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$ $\Delta H^\circ = -1000 \text{ kJ}$

2. Find standard enthalpy change for the formation of aluminum oxide and aluminum nitride.
 and: $\text{Al} = 0 \text{ kJ/mol}$ $\text{Al}_2\text{O}_3 = -1675.7 \text{ kJ/mol}$ $\text{AlN} = -317.6 \text{ kJ/mol}$
 $\text{O}_2 = 0 \text{ kJ/mol}$ $\text{N}_2 = 0 \text{ kJ/mol}$ $\text{Al}_2\text{O}_3 = -1675.7 \text{ kJ/mol}$
 Standard enthalpy formation: $\text{Al}_2\text{O}_3 = -1675.7 \text{ kJ/mol}$ $\text{AlN} = -317.6 \text{ kJ/mol}$

a. Reaction (a): $2\text{Al} + 3/2\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ $\Delta H^\circ = -1675.7 \text{ kJ}$
 or $2\text{Al} + 1.5\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ $\Delta H^\circ = -1675.7 \text{ kJ}$

b. Reaction: $2\text{Al} + 3/2\text{N}_2 \rightarrow 2\text{AlN}$ $\Delta H^\circ = -635.2 \text{ kJ}$
 or $2\text{Al} + 1.5\text{N}_2 \rightarrow 2\text{AlN}$ $\Delta H^\circ = -635.2 \text{ kJ}$

c. Add: $2\text{Al} + 3/2\text{O}_2 + 3/2\text{N}_2 \rightarrow 2\text{AlN} + \text{Al}_2\text{O}_3$ $\Delta H^\circ = -2810.9 \text{ kJ}$

3. Find Gibbs free energy for the combustion reaction of carbon monoxide and hydrogen, which was given in equation 10.16: $\text{CO} + \text{H}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ or 7).

Calculate the standard enthalpy change for this reaction from the following standard equations:

$\text{CO} + 1/2\text{O}_2 \rightarrow \text{CO}_2$ $\Delta H^\circ = -283 \text{ kJ}$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ $\Delta H^\circ = -571.6 \text{ kJ}$
 $\text{CO}_2 \rightarrow \text{CO} + 1/2\text{O}_2$ $\Delta H^\circ = +283 \text{ kJ}$

a. Write equation for eq. 10.16 and standard equation (a):
 $\text{CO} + \text{H}_2 \rightarrow \text{CO}_2$ $\Delta H^\circ = -283 \text{ kJ}$
 $\text{H}_2 + 1/2\text{O}_2 \rightarrow \text{H}_2\text{O}$ $\Delta H^\circ = -285.8 \text{ kJ}$
 or $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ $\Delta H^\circ = -571.6 \text{ kJ}$

b. Add:
 $\text{CO} + \text{H}_2 \rightarrow \text{CO}_2$ $\Delta H^\circ = -283 \text{ kJ}$
 $\text{H}_2 + 1/2\text{O}_2 \rightarrow \text{H}_2\text{O}$ $\Delta H^\circ = -285.8 \text{ kJ}$
 or $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ $\Delta H^\circ = -571.6 \text{ kJ}$

c. Add:
 $\text{CO} + \text{H}_2 \rightarrow \text{CO}_2$ $\Delta H^\circ = -283 \text{ kJ}$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ $\Delta H^\circ = -571.6 \text{ kJ}$

Hess Law Problems With Answers

Jin-Ying Zhang



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