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Chapter 2: Atoms

- 2.1** (a) NaClO_3 (b) AlF_3
- 2.2** (a) The mass number is $15 + 16 = 31$.
(b) The mass number is $86 + 136 = 222$.
- 2.3** (a) The element has 15 protons, making it phosphorus (P); its symbol is ${}_{15}^{31}\text{P}$.
(b) The element has 86 protons, making it radon (Rn); its symbol is ${}_{86}^{222}\text{Rn}$.
- 2.4** (a) The atomic number of mercury (Hg) is 80; that of lead (Pb) is 82.
(b) An atom of Hg has 80 protons; an atom of Pb has 82 protons.
(c) The mass number of this isotope of Hg is $80 + 120 = 200$; the mass number for this isotope of Pb is $82 + 120 = 202$.
(d) The symbols of these isotopes are ${}^{200}_{80}\text{Hg}$ and ${}^{202}_{82}\text{Pb}$.
- 2.5** The atomic number of iodine (I) is 53. The number of neutrons in each isotope is $125 - 53 = 72$ for iodine-125 and $131 - 53 = 78$ for iodine-131. The symbols for these two isotopes are ${}^{125}_{53}\text{I}$ and ${}^{131}_{53}\text{I}$.
- 2.6** The atomic weight is 6.941 amu, which is nearer to 7 amu than 6 amu. Therefore, lithium-7 is the more abundant isotope. The relative abundances for these two isotopes are 92.50 percent for lithium-7 and 7.50 percent for lithium-6.
- 2.7** This element has 13 electrons and, therefore, 13 protons. The element with atomic number 13 is Aluminum (Al).
- Ât :**
- 2.8** Both Democritus and Dalton believed that matter was composed of tiny indivisible particles referred to as atoms. The major difference between Democritus and Dalton is that Dalton based his theory on evidence rather than belief.
- 2.10** (a) Oxygen - an element (b) Table salt - a compound
(c) Sea water - a mixture (d) Wine - a mixture
(e) Air - a mixture (f) Silver - an element
(g) Diamond - an element (h) A pebble - a mixture
(i) Gasoline - a mixture (j) Milk - a mixture
(k) Carbon dioxide - a compound (l) Bronze - a mixture

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable
measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of
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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic
eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹ eg 2 HNO₃(aq) + KOH(aq) → KNO₃(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹

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