

12.2**FACTORS AFFECTING GAS PRESSURE****SECTION REVIEW****Objectives**

- Explain how the amount of gas and the volume of the container affect gas pressure.
- Infer the effect of temperature changes on the pressure exerted by a contained gas.

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms a chapter has introduced. In each sentence, insert the word or phrase which is most *likely* to make the sentence correct.

- As long as the temperature of an enclosed gas does not change, _____ 1. _____, doubling the number of gas particles _____ 2. _____ the pressure. When gas particles are removed from a container the pressure decreases. _____ 3. _____, halving the gas particles decreases the pressure by _____ 4. _____. When the number of gas particles remains constant, the pressure can be increased by _____ 5. _____ the volume. Reducing the volume of a contained gas by half _____ 6. _____ the pressure.
- When the number of gas particles and the volume of a gas remain constant, the pressure of a gas can be increased by increasing the _____ 7. _____. If the Kelvin temperature doubles, the pressure of the contained gas _____ 8. _____. If the pressure decreases by half when the Kelvin temperature is halved,

Part B True-False

Classify each of these statements as true always true, AT; sometimes true, ST; or never true, NT.

- _____ 9. Halving the number of particles in a given volume of gas decreases the pressure by one-half if the temperature is kept constant.
- _____ 10. Reducing the volume of a container of gas by one-half has the same effect on pressure as doubling the quantity of the gas in the container by one-half, if temperature is kept constant.
- _____ 11. When a sealed container of gas under pressure is opened, gas particles move from the container to the vacuum of air.

Gases And Pressure Section Review Answers

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Table of Contents Gases And Pressure Section Review Answers

1. Understanding the eBook Gases And Pressure Section Review Answers
 - The Rise of Digital Reading Gases And Pressure Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases And Pressure Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases And Pressure Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases And Pressure Section Review Answers
 - Personalized Recommendations
 - Gases And Pressure Section Review Answers User Reviews and Ratings
 - Gases And Pressure Section Review Answers and Bestseller Lists
5. Accessing Gases And Pressure Section Review Answers Free and Paid eBooks

- Gases And Pressure Section Review Answers Public Domain eBooks
 - Gases And Pressure Section Review Answers eBook Subscription Services
 - Gases And Pressure Section Review Answers Budget-Friendly Options
6. Navigating Gases And Pressure Section Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases And Pressure Section Review Answers Compatibility with Devices
 - Gases And Pressure Section Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases And Pressure Section Review Answers
 - Highlighting and Note-Taking Gases And Pressure Section Review Answers
 - Interactive Elements Gases And Pressure Section Review Answers
 8. Staying Engaged with Gases And Pressure Section Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases And Pressure Section Review Answers
 9. Balancing eBooks and Physical Books Gases And Pressure Section Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases And Pressure Section Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gases And Pressure Section Review Answers
 - Setting Reading Goals Gases And Pressure Section Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gases And Pressure Section Review Answers
 - Fact-Checking eBook Content of Gases And Pressure Section Review Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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