

Name: \_\_\_\_\_

# GAS LAWS

## Across

- For a gas, the volume that one mole occupies at 0 Celsius and 1 atm pressure.
- The law that states that for a fixed amount of gas at a constant temperature, the volume of the gas increases as the pressure on the gas decreases and the volume of the gas decreases as the pressure on the gas increases for identical masses of gas, at constant temperature.
- Process by which particles under pressure pass through a tiny opening.
- Units of  $R$  depend on units of measurement used for  $P$ ,  $V$  and  $TR = .0821$
- A gas that obeys none of the assumptions of the kinetic-molecular theory.
- State that equal volumes of gases at the same temperature and pressure contain equal numbers of particles.
- An imaginary gas that perfectly fits all the assumptions of the kinetic-molecular theory.



## Down

- Zero of the kinetic energy which represents the lowest possible theoretical temperature; atoms are all in the lowest possible energy state.
- An instrument that measures atmospheric pressure.
- The pressure of earth's atmosphere at sea level; exactly equivalent to 760 mm Hg.
- Gas particles collide with each other.
- The law that states that the pressure of a gas at a constant volume is directly proportional to the absolute temperature.
- A collision between gas particles and between gas particles and container walls in which there is no net loss of kinetic energy.
- The mathematical relationship among pressure, volume, temperature, and the number of moles of gas.
- The same as the tendency of a substance to move from a high concentration gradient from a more concentrated to a less concentrated area.
- The relationship between pressure, volume, and temperature of a fixed amount of gas.
- As Volume increases temperature increases if pressure is held constant.

# Gas Laws Vocabulary Chart Answers

**M Walker**



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