

Gas Laws and Theory Quiz SCH3U

Name: _____

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(Product) 7 7 4 6

Assessment of Learning

$P = 8.31 \text{ kPa} \cdot \text{L} / \text{mol} \cdot \text{K} \times 1 \text{ mol} = 101.325 \text{ kPa} = 760 \text{ torr} = 760 \text{ mmHg}$ SATP = 25°C and 100 kPa

$\text{Mol} \times R$
100.3 kPa

STP = 0°C and

Multiple Choice. Circle the most correct answer. (780)

1. A pressure reading of 103.4 kPa is equivalent to _____ mmHg.
A) 663.4 B) 775.6 C) 774.6 D) 780.4

2. The particles that make up a solid move _____ than do the particles that make up a gas.

A) more quickly and farther C) more quickly
B) more slowly D) in the same way

3. According to the kinetic-molecular theory, how does a gas expand?

A) Its particles become larger. B) Collisions between particles become elastic.
C) Its temperature decreases. D) Its particles move greater distances apart.

4. If the temperature of a container of gas remains constant, how could the pressure of the gas increase?

A) The mass of the gas molecules increase B) The diffusion of the gas molecules increases.
C) The size of the container increases D) The number of gas molecules in the container increases.

5. Which of the following laws shows an inverse relationship between pressure and volume?

A) Boyle's Law B) Charles's Law C) Graham's Law D) Dalton's Law

6. A gaseous mixture at a total pressure of 1.50 atm contains equal molar amounts of He, Ne, and Ar. At constant temperature CO_2 gas is added to the mixture until the total pressure is 3.00 atm. Which of the following is a correct statement concerning partial pressures after the CO_2 addition?

A) The partial pressure of Ar has doubled.
B) The partial pressure of CO_2 is 3 times that of Ne

Gas Laws Quiz Answers

**Thirumalaisamy P.
Velavancorresponding**



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