

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant  $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 \text{ kPa} / 1 \text{ atm}$  to get  
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^\circ\text{C}$ , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$\text{or } P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^\circ\text{C}$ , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of  $67^\circ\text{C}$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$\text{or } P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

# Gas Laws Practice Calculations Answers Unit Lesson 01

**Xiaolong Qi**



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## **Table of Contents Gas Laws Practice Calculations Answers Unit Lesson 01**

1. Understanding the eBook Gas Laws Practice Calculations Answers Unit Lesson 01
  - The Rise of Digital Reading Gas Laws Practice Calculations Answers Unit Lesson 01
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Calculations Answers Unit Lesson 01
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws Practice Calculations Answers Unit Lesson 01
  - User-Friendly Interface

4. Exploring eBook Recommendations from Gas Laws Practice Calculations Answers Unit Lesson 01
  - Personalized Recommendations
  - Gas Laws Practice Calculations Answers Unit Lesson 01 User Reviews and Ratings
  - Gas Laws Practice Calculations Answers Unit Lesson 01 and Bestseller Lists
5. Accessing Gas Laws Practice Calculations Answers Unit Lesson 01 Free and Paid eBooks
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Public Domain eBooks
  - Gas Laws Practice Calculations Answers Unit Lesson 01 eBook Subscription Services
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Budget-Friendly Options
6. Navigating Gas Laws Practice Calculations Answers Unit Lesson 01 eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Compatibility with Devices
  - Gas Laws Practice Calculations Answers Unit Lesson 01 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Laws Practice Calculations Answers Unit Lesson 01
  - Highlighting and Note-Taking Gas Laws Practice Calculations Answers Unit Lesson 01
  - Interactive Elements Gas Laws Practice Calculations Answers Unit Lesson 01
8. Staying Engaged with Gas Laws Practice Calculations Answers Unit Lesson 01
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Laws Practice Calculations Answers Unit Lesson 01
9. Balancing eBooks and Physical Books Gas Laws Practice Calculations Answers Unit Lesson 01
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Laws Practice Calculations Answers Unit Lesson 01
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Practice Calculations Answers Unit Lesson 01
  - Setting Reading Goals Gas Laws Practice Calculations Answers Unit Lesson 01
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Gas Laws Practice Calculations Answers Unit Lesson 01
  - Fact-Checking eBook Content of Gas Laws Practice Calculations Answers Unit Lesson 01
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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