

Mixed Gas Laws Worksheet

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$V = 98 \text{ L}$$

$$P = 2.8 \text{ atm}$$

$$T = 292 \text{ K}$$

$$n = ?$$

$$R = 0.082 \frac{\text{L atm}}{\text{mol K}}$$

$$PV = nRT$$

$$n = \frac{PV}{RT}$$

$$= \frac{2.8 \times 98}{0.082 \times 292} = \underline{\underline{11.46 \text{ mole}}}$$

- 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25°C , what will the pressure of the resulting mixture of gases be?

$$5 \text{ moles } \text{O}_2$$

$$3 \text{ moles } \text{N}_2$$

$$V = 30 \text{ L}$$

$$T = 25^\circ \text{C} \rightarrow 273 + 25 = \underline{\underline{298 \text{ K}}}$$

$$P = ?$$

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{8 \times 0.082 \times 298}{30 \text{ L}}$$

$$= \underline{\underline{6.5 \text{ atm}}}$$

- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C . By noon the temperature has risen to 45.0°C . What is the new volume of the balloon?

$$① V_1 = 35 \text{ L}$$

$$① T_1 = 20^\circ \text{C} = 293 \text{ K}$$

$$② T_2 = 45^\circ \text{C} = 318 \text{ K}$$

$$② V_2 = ?$$

Charles law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{35}{20} = \frac{V_2}{45}$$

$$V_2 = \frac{35 \times 318}{293} = \underline{\underline{37.78 \text{ L}}}$$

- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$V = 35 \text{ L}$$

$$T = 315 \text{ K}$$

$$P = 190 \text{ atm}$$

$$n = ?$$

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{190 \times 35}{0.082 \times 315}$$

$$= \underline{\underline{257.45 \text{ mol}}}$$

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