

Thermodynamics - Class Lecture Notes (Summary)

<p>1. A system is defined as a region of space containing matter. The boundary of the system is called the control surface.</p> <p>2. The system is classified into two types: open and closed.</p> <p>3. In an open system, both mass and energy can cross the control surface.</p> <p>4. In a closed system, mass cannot cross the control surface, but energy can.</p> <p>5. In an isolated system, neither mass nor energy can cross the control surface.</p> <p>6. The first law of thermodynamics states that the total energy of a system is conserved.</p> <p>7. The second law of thermodynamics states that the entropy of a system always increases.</p> <p>8. The third law of thermodynamics states that the entropy of a system approaches zero as the temperature approaches absolute zero.</p>	<p>9. The first law of thermodynamics can be written as:</p> $\Delta U = Q - W$ <p>where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.</p> <p>10. The second law of thermodynamics can be written as:</p> $\Delta S \geq 0$ <p>where ΔS is the change in entropy.</p> <p>11. The third law of thermodynamics can be written as:</p> $S \rightarrow 0 \text{ as } T \rightarrow 0$ <p>where S is the entropy and T is the temperature.</p>	<p>12. The first law of thermodynamics can be written as:</p> $\Delta U = Q - W$ <p>where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.</p> <p>13. The second law of thermodynamics can be written as:</p> $\Delta S \geq 0$ <p>where ΔS is the change in entropy.</p> <p>14. The third law of thermodynamics can be written as:</p> $S \rightarrow 0 \text{ as } T \rightarrow 0$ <p>where S is the entropy and T is the temperature.</p>
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