#### Gay-Lorenac's Law

- The gas in an serocol can is at a temperature of 27 °C and a pressure of 3 atm. If the gas reaches a
  pressure of 9 atm. what will the can's temperature be?
- 7. An oxygen tank reads 130 kPs with a temperature of 25 °C. What will the pressure read if the tank is cooled to -183 °C?

#### Combined Gas Law

- A gap has a volume of 35 libers at STP. What will its volume be at 5.0 atm and 35 °CP.
- Q.4.L of a gas is contained at .29 atm and 275 K. What will the volume of the gas be at .15 atm and 290KP.
- 400 mL of gas is stored at STP. The volume changes to 560 mL, and the temperature to 30 °C. What is the
  pressure lineates?

#### Ideal Gas Law

- 11. If 6 moles of a gas is at a pressure of 6.6 atm and a volume of 11 libers, what is the temperature of the gas?
- If there is unknown quantity of gas at a pressure of 123.6 kPs, a volume of 25 liters, and a temperature of 87 °C, how many mores of gas are in the sample?
- If there are 5 mores of gas in a 50-liter container, at a temperature of 250 K, what is the pressure of the gas within the container?
- 14. If 10.5 mores of gas are stored at a presourc of 0.15 atm and a temperature of 25 °C, what is the volume of the conceiner that the gas is store in?

#### Deficer's Law.

- 15. A storage container holds a mixture of oxygen, helium, and carbon dioxide. If the partial pressures of the three gases tank are 15 atm of O<sub>1</sub>, 8 atm of He, and 25 atm of OO<sub>2</sub>, what is the total pressure inside of the container?
- The total pressure of a gas miniture is 1.9 kPs. The partial pressure of gas A is 0.06 kPs and the partial pressure of gas B is 0.25 kPs. What is the partial pressure of the remaining gas in the miniture?
- If I place 30 moles of M, and 13 moles of CO, in a 40 L container at a temperature of 25° C, what will the pressure of the resulting mixture of asses be?

# **Gas Laws Guide Answer Key**

**AW Rasmussen** 

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