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Figure 1

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Gas Law Webquest Answers

Patrick Hurley



Gas Law Webquest Answers:

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Cases and Materials on Oil and Gas Law John S. Lowe, Owen L. Anderson, Ernest Edgar Smith, David E. Pierce, Christopher Stewart Kulander, 2013 Gas Laws Lifeliqe, 2019

This lesson plan covers Boyle's law, Charles' law, Gay Lussac's law, and Avogadro's law. The laws are applied to make various calculations concerning constant temperature, pressure, volume, and moles.

The Ideal Gas Law Handbook - Everything You Need to Know about Ideal Gas Law Patrick Hurley, 2016-04-29

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Van der Waals equation

Excerpt

The van der Waals equation is an equation of state for a fluid composed of particles that have a non-zero volume and a pairwise attractive inter-particle force, such as the van der Waals force. It was derived in 1873 by Johannes Diderik van der Waals, who received the Nobel prize in 1910 for his work on the equation of state for gases and liquids. The equation is based on a modification of the ideal gas law and approximates the behavior of real fluids, taking into account the non-zero size of molecules and the attraction between them. The van der Waals isotherms; the model correctly predicts a

mostly incompressible liquid phase but the oscillations in the phase transition zone do not fit experimental data The equation uses the following state variables the pressure of the fluid p total volume of the container containing the fluid V number of moles n and absolute temperature of the system T One form of the equation is where v is the volume of the container shared between each particle not the velocity of a particle N is the total number of particles and k is Boltzmann's constant given by the universal gas constant R and Avogadro's constant N_A Extra parameters are introduced a is a measure for the attraction between the particles and b is the average volume excluded from v by a particle The equation can be cast into the better known form where a is a measure of the attraction between the particles b is the volume excluded by a mole of particles A careful distinction

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