

Gas Laws Worksheet

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

Boyle's Law Problems:

$$P_1 V_1 = P_2 V_2$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mmHg})(22.5 \text{ L}) = (725 \text{ mmHg}) V_2$$

$$V_2 = \frac{(748 \text{ mmHg})(22.5 \text{ L})}{(725 \text{ mmHg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

Gas Law Problems And Answers

JA Banks



Gas Law Problems And Answers:

The book delves into Gas Law Problems And Answers. Gas Law Problems And Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Gas Law Problems And Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Gas Law Problems And Answers
 - Chapter 2: Essential Elements of Gas Law Problems And Answers
 - Chapter 3: Gas Law Problems And Answers in Everyday Life
 - Chapter 4: Gas Law Problems And Answers in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, the author will provide an overview of Gas Law Problems And Answers. This chapter will explore what Gas Law Problems And Answers is, why Gas Law Problems And Answers is vital, and how to effectively learn about Gas Law Problems And Answers.
 3. In chapter 2, this book will delve into the foundational concepts of Gas Law Problems And Answers. This chapter will elucidate the essential principles that must be understood to grasp Gas Law Problems And Answers in its entirety.
 4. In chapter 3, this book will examine the practical applications of Gas Law Problems And Answers in daily life. The third chapter will showcase real-world examples of how Gas Law Problems And Answers can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Gas Law Problems And Answers in specific contexts. The fourth chapter will explore how Gas Law Problems And Answers is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Gas Law Problems And Answers. The final chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Gas Law Problems And Answers.

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