# Gas Stoichiometry #1

If reactants and products are at the same conditions of temperature and pressure, then mole ratios of <u>gases</u> are also volume ratios.

$$3 H_2(g) + N_2(g) \rightarrow 2NH_3(g)$$

- 3 moles  $H_2$  + 1 mole  $N_2$   $\rightarrow$  2 moles  $NH_3$
- 3 liters  $H_2$  + 1 liter  $N_2$   $\rightarrow$  2 liters  $NH_3$

# **Gas Stoichiometry Edgenuity**

**AN Whitehead** 

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