

Worksheet – Gases II (Combined Gas Law, Ideal Gas Law, and Dalton's Law of Partial Pressure)

Combined Gas Law: $\frac{PV}{T} = \text{constant}$

1. A gas takes up a volume of 17 liters, has a pressure of 2.3 atm, and a temperature of 299 K. If I raise the temperature to 350 K and lower the pressure to 1.5 atm, what is the new volume of the gas?

amount of gas is fixed. so combined gas law. $\frac{PV}{T} = \text{const.}$

$P_1 = 2.3 \text{ atm}$ $P_2 = 1.5 \text{ atm}$

$V_1 = 17 \text{ L}$ $V_2 = ? \text{ L}$

$T_1 = 299 \text{ K}$ $T_2 = 350 \text{ K}$

$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$

$V_2 = \frac{P_1 V_1}{T_1} \cdot \frac{T_2}{P_2} = \frac{2.3 \text{ atm} \times 17 \text{ L}}{299 \text{ K}} \times \frac{350 \text{ K}}{1.5 \text{ atm}} = 30.5 \text{ L}$

2. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50°C, what will be the temperature of the gas if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

same approach.

$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$

$P_1 = 5 \text{ atm}$

$P_2 = 3 \text{ atm}$

$V_1 = 2.9 \text{ L}$

$V_2 = 2.4 \text{ L}$

$T_1 = 50^\circ\text{C}$
 $= 323 \text{ K}$

$T_2 = ? \text{ K}$

$T_2 = \frac{P_2 V_2}{P_1 V_1} T_1 = \frac{3 \text{ atm} \times 2.4 \text{ L}}{5 \text{ atm} \times 2.9 \text{ L}} \times 323 \text{ K} = 160 \text{ K}$

3. I have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 liters, how many liters of gas did I start with?

same approach.

$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$

$P_1 = 60 \text{ atm}$

$P_2 = 30 \text{ atm}$

$V_1 = ? \text{ L}$

$V_2 = 29 \text{ L}$

$T_1 = 115 \text{ K}$

$T_2 = 225 \text{ K}$

$V_1 = \frac{P_2 V_2}{T_2} \times \frac{T_1}{P_1} = \frac{30 \text{ atm} \times 29 \text{ L} \times 115 \text{ K}}{225 \text{ K} \times 60 \text{ atm}} = 7.4 \text{ L}$

4. If I have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of the gas if I decrease the pressure to 45 atm and decrease the temperature to 750 K? [30.31]

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