

9

Thermodynamics

- Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
 - ΔS is negative and therefore, ΔH should be highly positive
 - ΔS is negative and therefore, ΔH should be highly negative
 - ΔS is positive and therefore, ΔH should be negative
 - ΔS is positive and therefore, ΔH should also be highly positive
- For the reaction, $X_2O_4(l) \longrightarrow 2XO_2(g)$, $\Delta U = 2.1 \text{ kcal}$, $\Delta S = 20 \text{ cal K}^{-1}$ at 300 K. Hence, ΔG is [2014]
 - 2.7 kcal
 - 2.7 kcal
 - 9.3 kcal
 - 9.3 kcal
- For a given exothermic reaction, K_p and K_p' are the equilibrium constants at temperatures T_1 and T_2 , respectively. Assuming that heat of reaction is constant in temperature range between T_1 and T_2 , it is readily observed that [2014]
 - $K_p > K_p'$
 - $K_p < K_p'$
 - $K_p = K_p'$
 - $K_p = \frac{1}{K_p'}$
- A reaction having equal energies of activation for forward and reverse reactions has [2013]
 - $\Delta S = 0$
 - $\Delta G = 0$
 - $\Delta H = 0$
 - $\Delta H = \Delta G = \Delta S = 0$
- In which of the following reactions, standard reaction entropy changes (ΔS°) is positive and standard Gibbs energy change (ΔG°) decreases sharply with increasing temperature? [2012]
 - $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
 - $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
 - $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
 - $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
- The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
 - 10.52 cal/mol K
 - 21.04 cal/mol K
 - 5.260 cal/mol K
 - 0.526 cal/mol K
- Standard enthalpy of vaporisation $\Delta_{\text{vap}}H^\circ$ for water at 100°C is 40.66 kJ mol⁻¹. The internal energy of vaporisation of water at 100°C (in kJ mol⁻¹) is (assume water vapour to behave like an ideal gas). [2012]
 - + 37.56
 - 43.76
 - + 43.76
 - + 40.66
- If the enthalpy change for the transition of liquid water to steam is 30 kJ mol⁻¹ at 27°C, the entropy change for the process would be [2011]
 - 1.0 J mol⁻¹K⁻¹
 - 0.1 J mol⁻¹K⁻¹
 - 100 J mol⁻¹K⁻¹
 - 10 J mol⁻¹K⁻¹

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A TEXTBOOK OF CHEMICAL ENGINEERING THERMODYNAMICS K. V. NARAYANAN, 2013-01-11 Designed as an undergraduate level textbook in Chemical Engineering this student friendly thoroughly class room tested book now in its second edition continues to provide an in depth analysis of chemical engineering thermodynamics The book has been so organized that it gives comprehensive coverage of basic concepts and applications of the laws of thermodynamics in the initial chapters while the later chapters focus at length on important areas of study falling under the realm of chemical thermodynamics The reader is thus introduced to a thorough analysis of the fundamental laws of thermodynamics as well as their applications to practical situations This is followed by a detailed discussion on relationships among thermodynamic properties and an exhaustive treatment on the thermodynamic properties of solutions The role of phase equilibrium thermodynamics in design analysis and operation of chemical separation methods is also deftly dealt with Finally the chemical reaction equilibria are skillfully explained Besides numerous illustrations the book contains over 200 worked examples over 400 exercise problems all with answers and several objective type questions which enable students to gain an in depth understanding of the concepts and theory discussed The book will also be a useful text for students pursuing courses in chemical engineering related branches such as polymer engineering petroleum engineering and safety and environmental engineering New to This Edition More Example Problems and Exercise Questions in each chapter Updated section on Vapour Liquid Equilibrium in Chapter 8 to highlight the significance of equations of state approach GATE Questions up to 2012 with answers [Engineering Physics MCQ \(Multiple Choice Questions\)](#) Arshad Iqbal, The Engineering Physics Multiple Choice Questions MCQ Quiz with Answers PDF Engineering Physics MCQ PDF Download Quiz Questions Chapter 1 36 Practice Tests with Answer Key Physics Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs Engineering Physics MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests Engineering Physics MCQ PDF book helps to practice test questions from exam prep notes The Engineering Physics MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs Engineering Physics Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters Alternating fields and currents astronomical data capacitors and capacitance circuit theory conservation of energy coulomb s law current produced magnetic field electric potential energy equilibrium indeterminate structures finding electric field first law of thermodynamics fluid statics and dynamics friction drag and centripetal force fundamental constants of physics geometric optics inductance kinetic energy longitudinal waves magnetic force models of magnetism newton s law of motion Newtonian gravitation Ohm s law optical diffraction optical interference physics and measurement properties of common elements rotational motion second law of thermodynamics simple harmonic motion special relativity straight line motion transverse waves two and three dimensional motion vector quantities work

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Table of Contents Engineering Thermodynamics Test Questions

1. Understanding the eBook Engineering Thermodynamics Test Questions
 - The Rise of Digital Reading Engineering Thermodynamics Test Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Engineering Thermodynamics Test Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Engineering Thermodynamics Test Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Engineering Thermodynamics Test Questions
 - Personalized Recommendations
 - Engineering Thermodynamics Test Questions User Reviews and Ratings
 - Engineering Thermodynamics Test Questions and Bestseller Lists
5. Accessing Engineering Thermodynamics Test Questions Free and Paid eBooks
 - Engineering Thermodynamics Test Questions Public Domain eBooks
 - Engineering Thermodynamics Test Questions eBook Subscription Services
 - Engineering Thermodynamics Test Questions Budget-Friendly Options

6. Navigating Engineering Thermodynamics Test Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Engineering Thermodynamics Test Questions Compatibility with Devices
 - Engineering Thermodynamics Test Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Engineering Thermodynamics Test Questions
 - Highlighting and Note-Taking Engineering Thermodynamics Test Questions
 - Interactive Elements Engineering Thermodynamics Test Questions
8. Staying Engaged with Engineering Thermodynamics Test Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Engineering Thermodynamics Test Questions
9. Balancing eBooks and Physical Books Engineering Thermodynamics Test Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Engineering Thermodynamics Test Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Engineering Thermodynamics Test Questions
 - Setting Reading Goals Engineering Thermodynamics Test Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Engineering Thermodynamics Test Questions
 - Fact-Checking eBook Content of Engineering Thermodynamics Test Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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