

- (b) (i) Based on Maxwell's relations, prove that for any fluid,
- $$dS = C_v \frac{dT}{T} + \left( \frac{\partial p}{\partial T} \right)_v dV. \quad (8)$$

- (ii) A mixture of 3 moles of helium, 4 moles of neon and 5 moles of argon is at 1 bar and 300°K. Calculate

- (1) volume,
- (2) mole fraction and partial pressure of gases and
- (3) change of entropy due to mixing. (8)

15. (a) The moist air is at 45°C Dry Bulb Temperature and 30°C Wet Bulb Temperature. Calculate

- (i) Vapour pressure
- (ii) Dew point temperature
- (iii) Specific enthalpy
- (iv) Relative humidity
- (v) Degree of saturation
- (vi) Vapour density
- (vii) Enthalpy of mixture. (16)

Or

- (b) (i) Explain the process of cooling with dehumidification of air, with the help of sketches of schematic of air duct with the elements involved as well as the representation on psychrometric chart. (6)

- (ii) An air-water vapour mixture at 20°C and 50% relative humidity at a pressure of 1.013 bar is heated at constant pressure to a temperature of 35°C. Calculate

- (1) The initial and final specific humidity
- (2) Final relative humidity
- (3) Dew point temperature
- (4) Heat transferred per kg of dry air. (10)

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