

5**ELECTRONS IN ATOMS****Practice Problems***In your notebook, solve the following problems.***SECTION 5.1 MODELS OF THE ATOM**

- How many sublevels are in the following principal energy levels?
 - $n = 1$
 - $n = 2$
 - $n = 3$
 - $n = 4$
 - $n = 5$
 - $n = 6$
- How many orbitals are in the following sublevels?
 - 1s sublevel
 - 5s sublevel
 - 4d sublevel
 - 4f sublevel
 - 7s sublevel
 - 3p sublevel
 - fifth principal energy level
 - 6d sublevel
- What are the types of sublevels and number of orbitals in the following energy levels?
 - $n = 1$
 - $n = 2$
 - $n = 3$
 - $n = 4$
 - $n = 5$

SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS

- Write a complete electron configuration of each atom.
 - hydrogen
 - vanadium
 - magnesium
 - barium
 - bromine
 - sulfur
 - krypton
 - arsenic
 - radon

SECTION 5.3 PHYSICS AND THE QUANTUM MECHANICAL MODEL

- What is the wavelength of the radiation whose frequency is $5.00 \times 10^{15} \text{ s}^{-1}$? In what region of the electromagnetic spectrum is this radiation?
- An inexpensive laser that is available to the public emits light that has a wavelength of 670 nm. What are the color and frequency of the radiation?
- What is the energy of a photon whose frequency is $2.22 \times 10^{14} \text{ s}^{-1}$?
- What is the frequency of a photon whose energy is $6.00 \times 10^{-19} \text{ J}$?
- Arrange the following types of electromagnetic radiation in order of increasing frequency.
 - infrared
 - cosmic rays
 - visible light
 - radio waves
 - microwaves
 - ultraviolet
- Suppose that your favorite AM radio station broadcasts at a frequency of 1600 kHz. What is the wavelength in meters of the radiation from the station?

Electrons In Atoms Answer Key

Arthur Winter



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