

Chapter 21 • Electrochemistry

(See page 3 for formulas and an answer key)

PRACTICE TEST

1. Which of the following is the correct cell notation for the reaction



- $\text{Cd}^{2+} | \text{Cd} || \text{Hg}_2^{2+} | \text{Hg}$
- $\text{Cd}^{2+} | \text{Hg}_2^{2+} || \text{Cd} | \text{Hg}$
- $\text{Cd} | \text{Cd}^{2+} || \text{Hg}_2^{2+} | \text{Hg}$
- $\text{Cd}^{2+} | \text{Hg} || \text{Hg}_2^{2+} | \text{Cd}$
- $\text{Hg} | \text{Cd} || \text{Hg}_2^{2+} | \text{Cd}^{2+}$

2. Consider an electrochemical cell where the following reaction takes place:



Which of the following is the correct cell notation for this cell?

- $\text{Al} | \text{Al}^{3+} || \text{Sn}^{2+} | \text{Sn}$
- $\text{Al}^{3+} | \text{Al} || \text{Sn} | \text{Sn}^{2+}$
- $\text{Sn} | \text{Sn}^{2+} || \text{Al}^{3+} | \text{Al}$
- $\text{Sn} | \text{Al}^{3+} || \text{Al} | \text{Sn}^{2+}$
- $\text{Al} | \text{Sn}^{2+} || \text{Sn} | \text{Al}^{3+}$

Standard Reduction Potentials at 25°C E° (volts)

$\text{F}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{F}^-(\text{aq})$	+2.87
$\text{Au}^{3+} + 3\text{e}^- \rightarrow \text{Au(s)}$	+1.50
$\text{Cl}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{Cl}^-(\text{aq})$	+1.36
$\text{O}_2(\text{g}) + 4\text{H}_3\text{O}^+(\text{aq}) + 4\text{e}^- \rightarrow 6\text{H}_2\text{O(l)}$	+1.23
$\text{Br}_2(\text{l}) + 2\text{e}^- \rightarrow 2\text{Br}^-(\text{aq})$	+1.08
$\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag(s)}$	+0.80
$\text{Hg}_2^{2+}(\text{aq}) + 2\text{e}^- \rightarrow 2\text{Hg(l)}$	+0.79
$\text{I}_2(\text{s}) + 2\text{e}^- \rightarrow 2\text{I}^-(\text{aq})$	+0.535
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu(s)}$	+0.337
$\text{Sn}^{4+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn}^{2+}(\text{aq})$	+0.15
$\text{Sn}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn(s)}$	-0.14
$\text{Cd}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cd(s)}$	-0.40
$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Zn(s)}$	-0.763
$2\text{H}_2\text{O(l)} + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$	-0.828
$\text{Al}^{3+}(\text{aq}) + 3\text{e}^- \rightarrow \text{Al(s)}$	-1.66
$\text{K}^+(\text{aq}) + \text{e}^- \rightarrow \text{K(s)}$	-2.93
$\text{Li}^+(\text{aq}) + \text{e}^- \rightarrow \text{Li(s)}$	-3.045

3. Given the two half reactions and their potentials, which net reaction is spontaneous?



- $\text{Ni(s)} + \text{Mg}^{2+}(\text{aq}) \rightarrow \text{Mg(s)} + \text{Ni}^{2+}(\text{aq})$
- $\text{Ni}^{2+}(\text{aq}) + \text{Mg(s)} \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{Ni(s)}$
- $\text{Ni(s)} + \text{Mg(s)} \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{Ni}^{2+}(\text{aq})$
- $\text{Mg}^{2+}(\text{aq}) + \text{Ni}^{2+}(\text{aq}) \rightarrow \text{Mg(s)} + \text{Ni(s)}$
- $\text{Mg}^{2+}(\text{aq}) + \text{Mg(s)} \rightarrow \text{Ni(s)} + \text{Ni}^{2+}(\text{aq})$

4. Calculate E° for the following reaction:



- +6.00 V
- 3.08 V
- +3.08 V
- +2.78 V
- 2.78 V

5. Calculate E° for the following reaction:



- 2.06 V
- +4.52 V
- +2.06 V
- 4.52 V
- 1.26 V

6. Using data from the reduction potential table and the reaction



calculate the standard reduction potential of the half-reaction



- 1.18 V
- 0.40 V
- 0.40 V
- 1.18 V
- 2.00 V

7. Using data from the reduction potential table, predict which of the following is the best oxidizing agent.

- F_2
- Ag
- Sn^{4+}
- Ag^+
- Al^{3+}

Electrochemistry 21 Chapter Test A Answer Key

**Peter Horsman, Brian E. Conway, E.
Yeager**



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Jespersen, Pamela Kerrigan, 2023-07-04 Always study with the most up to date prep Look for AP Chemistry Premium 2025 Prep Book with 6 Practice Tests Comprehensive Review Online Practice ISBN 9781506291802 on sale July 2 2024 Publisher's Note Products purchased from third party sellers are not guaranteed by the publisher for quality authenticity or access to any online entities included with the product Handbook of Electrochemistry Cynthia G. Zoski, 2006-12-11

Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical and biological systems corrosion energy applications involving fuel cells and solar cells and nanoscale investigations The Handbook of Electrochemistry serves as a source of electrochemical information providing details of experimental considerations representative calculations and illustrations of the possibilities available in electrochemical experimentation The book is divided into five parts Fundamentals Laboratory Practical Techniques Applications and Data The first section covers the fundamentals of electrochemistry which are essential for everyone working in the field presenting an overview of electrochemical conventions terminology fundamental equations and electrochemical cells experiments literature textbooks and specialized books Part 2 focuses on the different laboratory aspects of electrochemistry which is followed by a review of the various electrochemical techniques ranging from classical experiments to scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry Applications of electrochemistry include electrode kinetic determinations unique aspects of metal deposition and electrochemistry in small places and at novel interfaces and these are detailed in Part 4 The remaining three chapters provide useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials serves as a source of electrochemical information includes useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials reviews electrochemical techniques incl scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry

Principles and Applications of Electrochemistry Henry Jermain Maude Creighton, 1924 **Solved Problems in**

Electrochemistry for Universities and Industry Dominique L. Piron, 2004 **Electro Chemistry**, SAT Subject Test:

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Volume 1 Modern Electrochemistry John O'M. Bockris, Amulya K. N. Reddy, 2013-11-21. This book had its nucleus in some lectures given by one of us, J. O. M. B., in a course on electrochemistry to students of energy conversion at the University of Pennsylvania. It was there that he met a number of people trained in chemistry, physics, biology, metallurgy, and materials science, all of whom wanted to know something about electrochemistry. The concept of writing a book about electrochemistry, which could be understood by people with very varied backgrounds, was thereby engendered. The lectures

were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Corrosion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States

Modern Electrochemistry John Bockris, 2012-12-06 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Corrosion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States

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2A John O'M. Bockris, Amulya K.N. Reddy, Maria E. Gamboa-Aldeco, 2001-01-31 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Corrosion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States Comprehensive Treatise of Electrochemistry Peter Horsman, Brian E. Conway, E. Yeager, 2013-11-11 It is now time for a comprehensive treatise to look at the whole field of electrochemistry The present treatise was conceived in 1974 and the earliest invitations to authors for contributions were made in 1975 The completion of the early volumes has been delayed by various factors There has been no attempt to make each article emphasize the most recent situation at the expense of an overall statement of the modern view This treatise is not a collection of articles from Recent Advances in Electrochemistry or Modern Aspects of Electrochemistry It is an attempt at making a mature statement about the present position in the vast area of what is best looked at as a new interdisciplinary field Texas A M University J O M Bockris University of Ottawa B E Conway Case Western Reserve University Ernest Yeager Texas A M University Ralph E White Preface to Volume 3 Of events

which have affected progress in the field of electrochemistry the decision of NASA to use electrochemical auxiliary power in space vehicles was one of the more important Another important decision was Ford's announcement of their sodium sulfur cell for vehicular use in 1969

Surface Electrochemistry John O'M. Bockris, Shahad U.M. Khan, 2013-03-07 The text Modern Electrochemistry authored by J O M Bockris and A K N Reddy and published by Plenum Press in 1970 was written between 1967 and 1969 The concept for it arose in 1962 in the Energy Conversion Center at the University of Pennsylvania and it was intended to act as a base for interdisciplinary students and mature scientists chemists physicists biologists metallurgists and engineers who wanted to know about electrochemical energy conversion and storage In writing the book the stress therefore was placed above all on lucidity in teaching physical electrochemistry from the beginning Although this fundamentally undergraduate text continues to find purchasers 20 years after its birth it has long been clear that a modernized edition should be written and the plans to do so were the origin of the present book However if a new Bockris and Reddy was to be prepared and include the advances of the last 20 years with the same degree of lucidity as characterized the first one the depth of the development would have to be well short of that needed by professional electrochemists

Encyclopedia of Surface and Colloid Science, 2004 Update Supplement P. Somasundaran, 2004-08-11 Appending the Encyclopedia of Surface and Colloid Science by 42 entries as well as 3800 new citations 1012 equations and 485 illustrations and chemical structures this important supplement summarizes a constellation of new theoretical and experimental findings related to chemical characterization mechanisms interfacial behavior methods and modeling and applications

Introduction to Corrosion Science Edward McCafferty, 2010-07-17 This textbook is intended for a one semester course in corrosion science at the graduate or advanced undergraduate level The approach is that of a physical chemist or materials scientist and the text is geared toward students of chemistry materials science and engineering This textbook should also be useful to practicing corrosion engineers or materials engineers who wish to enhance their understanding of the fundamental principles of corrosion science It is assumed that the student or reader does not have a background in electrochemistry However the student or reader should have taken at least an undergraduate course in materials science or physical chemistry More material is presented in the textbook than can be covered in a one semester course so the book is intended for both the classroom and as a source book for further use This book grew out of classroom lectures which the author presented between 1982 and the present while a professorial lecturer at George Washington University Washington DC where he organized and taught a graduate course on Environmental Effects on Materials Additional material has been provided by over 30 years of experience in corrosion research largely at the Naval Research Laboratory Washington DC and also at the Bethlehem Steel Company Bethlehem PA and as a Robert A Welch Postdoctoral Fellow at the University of Texas The text emphasizes basic principles of corrosion science which underpin extensions to practice

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The Porphyrin Handbook, Volume 1 Karl Kadish, Kevin M. Smith, Roger Guilard, 2000 Scientists in such fields as mathematics physics chemistry biochemistry biology and medicine are currently involved in investigations of porphyrins and their numerous analogues and derivatives Porphyrins are being used as platforms for the study of theoretical principles as catalysts as drugs as electronic devices and as spectroscopic probes in biology and medicine The need for an up to date and authoritative treatise on the porphyrin system has met with universal acclaim amongst scientists and investigators

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Analytical Electrochemistry Joseph Wang, 2006-04-27 Third Edition covers the latest advances in methodologies sensors detectors and mIcrochips The greatly expanded Third Edition of this internationally respected text continues to provide readers with a complete panorama of

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