

10 • Orbital Hybridization & Molecular Orbitals

PRACTICE TEST

1. Which hybridization is associated with a steric number of 3?

- a) sp 2
b) sp^2 3
 c) sp^3 4

- d) sp^3d 5
 e) sp^3d^2 6

2. The molecule BrF_3 has a steric number of 5 on the central atom?

- a) 3 b) 4 **c) 5** d) 6



3. What is the hybridization of Br in BrF_3 ?

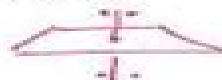
- a) sp
b) sp^2
 c) sp^3
d) sp^3d SN=5
 e) sp^3d^2

4. How many equivalent sp^3d orbitals are there?

- a) 3 b) 1 **c) 5** d) 6

5. What type of hybridization is associated with a square planar molecular shape?

- a) sp^3
 b) sp^2
 c) sp
d) sp^3d
e) sp^3d^2



6. What shape for electron pairs is associated with sp^3d^2 hybridization?

- a) linear d) tetrahedral
 b) square planar **c) octahedral**
 c) bent



7. What hybridization is predicted for phosphorus in the PCl_3 molecule?

- a) sp^2
b) sp^3
 c) sp
 d) sp^3d^2



tetrahedral shape = pyramidal

8. A double bond contains 1 sigma bond(s) and 1 pi bond(s).

- a) 0, 2 b) 1, 2 c) 2, 0 **d) 1, 1**

9. What angle exists between orbitals in sp^3d^2 hybrid orbitals?

- a) 90.0°** d) 120.0°
 b) 180.0° e) 78.5°
 c) 109.5°



10. Which of the following elements is most likely to display sp^3d hybridization? SN=5

- a) oxygen d) carbon
 b) nitrogen e) boron
c) phosphorus } no d orbitals

11. How many sigma (σ) and pi (π) electron pairs are in a carbon dioxide molecule? O=C=O

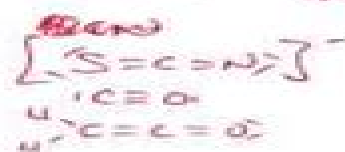
- a) four σ and zero π d) two σ and four π
 b) three σ and two π e) one σ and three π
c) two σ and two π

12. What is the hybridization of the oxygen atoms in CH_3OH and CO_2 , respectively? O=C=O

- a) sp^3 , sp^3 d) sp^2 , sp^2 SN=3
 b) sp^3 , sp^2 e) sp^3 , sp sp^2
c) sp^2 , sp^2

13. All of the following species contain two π -bonds EXCEPT

- a) SCN^- 16e- d) OCS 16e-
 b) CO 10e- **e) NO^+ 12e-**
 c) H_2CCO 16e-



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