

## Stoichiometry Worksheet 2: Percent Yield

What parts of a stoichiometry problem:

- Identify a balanced chemical equation.
- Calculate what you want to be produced, what amount given, what you have to work up.
- Identify and use units and/or chemical equations (grams, moles, liters, mass, mass, mass).

1. Using the following example as a guide, identify the given, what you want to be produced, and the units of each. Then, use stoichiometry to calculate the amount of product you can produce from the given amount of reactant.

React:  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$

Reactant:  $2 \text{ mol}$   $2 \text{ mol}$   $1 \text{ mol}$

Product:  $2 \text{ mol}$   $1 \text{ mol}$

Given:  $2 \text{ mol}$   $1 \text{ mol}$

$$\frac{2 \text{ mol } \text{H}_2\text{O}_2}{1 \text{ mol } \text{H}_2\text{O}_2} \times \frac{1 \text{ mol } \text{O}_2}{2 \text{ mol } \text{H}_2\text{O}_2} = 1 \text{ mol } \text{O}_2$$

$$1 \text{ mol } \text{O}_2 \times \frac{32 \text{ g } \text{O}_2}{1 \text{ mol } \text{O}_2} = 32 \text{ g } \text{O}_2$$

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**J Rink**



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### **Table of Contents Modeling Chemistry U8 Ws 1 V2 Key**

1. Understanding the eBook Modeling Chemistry U8 Ws 1 V2 Key
  - The Rise of Digital Reading Modeling Chemistry U8 Ws 1 V2 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U8 Ws 1 V2 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U8 Ws 1 V2 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U8 Ws 1 V2 Key
  - Personalized Recommendations
  - Modeling Chemistry U8 Ws 1 V2 Key User Reviews and Ratings
  - Modeling Chemistry U8 Ws 1 V2 Key and Bestseller Lists

5. Accessing Modeling Chemistry U8 Ws 1 V2 Key Free and Paid eBooks
  - Modeling Chemistry U8 Ws 1 V2 Key Public Domain eBooks
  - Modeling Chemistry U8 Ws 1 V2 Key eBook Subscription Services
  - Modeling Chemistry U8 Ws 1 V2 Key Budget-Friendly Options
6. Navigating Modeling Chemistry U8 Ws 1 V2 Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry U8 Ws 1 V2 Key Compatibility with Devices
  - Modeling Chemistry U8 Ws 1 V2 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U8 Ws 1 V2 Key
  - Highlighting and Note-Taking Modeling Chemistry U8 Ws 1 V2 Key
  - Interactive Elements Modeling Chemistry U8 Ws 1 V2 Key
8. Staying Engaged with Modeling Chemistry U8 Ws 1 V2 Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry U8 Ws 1 V2 Key
9. Balancing eBooks and Physical Books Modeling Chemistry U8 Ws 1 V2 Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry U8 Ws 1 V2 Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U8 Ws 1 V2 Key
  - Setting Reading Goals Modeling Chemistry U8 Ws 1 V2 Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U8 Ws 1 V2 Key
  - Fact-Checking eBook Content of Modeling Chemistry U8 Ws 1 V2 Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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