

Unit 6 – Worksheet 3

Ionic Compounds

Properties

Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- a. Na_2O
- b. K_2S
- c. MgCl_2
- d. CaBr_2
- e. BaI_2
- f. Al_2S_3
- g. CsBr
- h. AgF

2. Give the name of the following simple binary ionic compounds.

- a. Na_3N
- b. K_2O
- c. AgBr
- d. MgI_2
- e. SrO

3. Write the formula for the following binary ionic compounds.

- a. lithium bromide
- b. sodium iodide
- c. silver sulfide
- d. cesium oxide
- d. beryllium iodide
- f. barium hydride
- g. aluminum fluoride
- h. potassium oxide

4. Write the formula for these ionic substances.

- a. silver oxide
- b. aluminum sulfide
- c. sodium nitride
- d. barium chloride
- e. strontium hydride

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Karin Nielsen-Saines



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