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Chemistry – Unit 10 Worksheet 3

1. Element X has two natural isotopes. The isotope with a mass number of 6 has a relative abundance of 7.5%. The isotope with a mass number of 7 has a relative abundance of 92.5%. Determine the average molar mass for the element from these figures. What is the true identity and atomic number of element X?
2. The element copper is found to contain the naturally occurring isotopes $^{63}_{29}\text{Cu}$ and $^{65}_{29}\text{Cu}$. The relative abundances are 69.1% and 30.9% respectively. Calculate the average molar mass of copper.
3. Uranium has three isotopes with the following percent abundances: $^{234}_{92}\text{U}$ (0.0058%), $^{235}_{92}\text{U}$ (0.71%), $^{238}_{92}\text{U}$ (99.23%). What do you expect the molar mass of uranium to be in whole numbers? Why?
4. A sample of silver as it occurs in nature is 52.0% of isotope $^{107}_{47}\text{Ag}$ and 48.0% of isotope $^{109}_{47}\text{Ag}$. What is the average molar mass of silver? (Compare your result with the value given in the periodic table).

Modeling Chemistry Unit 1

Ensheng Dong



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Table of Contents Modeling Chemistry Unit 1

1. Understanding the eBook Modeling Chemistry Unit 1
 - The Rise of Digital Reading Modeling Chemistry Unit 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 1
 - Personalized Recommendations
 - Modeling Chemistry Unit 1 User Reviews and Ratings
 - Modeling Chemistry Unit 1 and Bestseller Lists
5. Accessing Modeling Chemistry Unit 1 Free and Paid eBooks
 - Modeling Chemistry Unit 1 Public Domain eBooks
 - Modeling Chemistry Unit 1 eBook Subscription Services
 - Modeling Chemistry Unit 1 Budget-Friendly Options
6. Navigating Modeling Chemistry Unit 1 eBook Formats

- ePub, PDF, MOBI, and More
- Modeling Chemistry Unit 1 Compatibility with Devices
- Modeling Chemistry Unit 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 1
 - Highlighting and Note-Taking Modeling Chemistry Unit 1
 - Interactive Elements Modeling Chemistry Unit 1
- 8. Staying Engaged with Modeling Chemistry Unit 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 1
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 1
 - Setting Reading Goals Modeling Chemistry Unit 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 1
 - Fact-Checking eBook Content of Modeling Chemistry Unit 1
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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