

Unit 2 Worksheet 1: Molar relationships

For each of the problems below:

- Write a full balanced chemical equation.
- Calculate, where it applies, molar masses and select the most useful molar ratio.
- How much of a mass balanced experimental demonstration exists?
- How can you be certain it is?

1. Hydrogen and nitrogen, which react to form nitrogen gas, form an ammonia gas (which is also a gas) and water. How many moles of nitrogen gas would be needed to completely form 1 mole of hydrogen gas?

Equation: $\text{H}_2 + \text{N}_2 \rightarrow \text{NH}_3$

Molar: 1 2 1 1

Change: 1.0 g 1.0 g 1.0 g 1.0 g

Answer: $\text{H}_2 + \text{N}_2 \rightarrow \text{NH}_3$

2. Propane, C_3H_8 , reacts to give CO_2 and H_2O . How many moles of CO_2 would be formed from 1.0 g of propane (molar mass 44.1 g/mol)?

Equation: $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

Molar: 1 5 3 4

Change: 1.0 g 1.0 g 1.0 g 1.0 g

Answer: $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

3. Ammonia, NH_3 , the nitrogen is made by reacting hydrogen and nitrogen to form an ammonia gas (which is a gas). How many moles of a nitrogen gas would be needed to form 1.0 g of ammonia gas?

Equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$

Molar: 1 3 2

Change: 1.0 g 1.0 g 1.0 g

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