Name		
	T223238	1553

## Honors Chemistry - Unit 8 Worksheet 1 Mole relationships

For each of the problems below:

- a. Write the balanced chemical equation
- b. Identify what is given (with units) and what you want to find (with units)
- Use coefficients from balanced equation to determine mole ratio
- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 2.0 moles of hydrogen sulfide?

Equation:  $_{-}$  H<sub>2</sub>S<sub> $\omega$ </sub> +  $_{-}$  O<sub>2</sub> $_{\omega}$   $\rightarrow$   $_{-}$  SO<sub>2</sub> $_{\omega}$  +  $_{-}$  H<sub>2</sub>O $_{\omega}$ 

Before:

Change.

 Aluminum reacts with hydrochloric acid to produce aluminum chloride and hydrogen gas. Write a balanced equation for the reaction. Calculate the number of moles of hydrochloric acid required to react with 0.50 moles of Al.

Equation:

Biefore:

Change, After

 Ammonia, NH3, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 1.25 moles of nitrogen gas?

Equation:

Biefore:

Change

Ather

# **Modeling Chemistry U8 Vs 2 0 Chemistryquoz**

**D** Kirk

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