

Chemistry – Life & Physical Sciences 1 – Portfolio March

Unit 1: Molecular Structure

Use Lewis structures to determine the molecular geometry of phosphorus trihydride and phosphorus pentachloride.

- Draw a Lewis structure of phosphorus trihydride and determine whether it has a trigonal pyramidal geometry.
- Draw a Lewis structure of phosphorus pentachloride and determine whether it has a trigonal bipyramidal geometry.

Draw Lewis structures for the following molecules and determine their molecular geometry. Use VSEPR to determine the molecular geometry.

- Phosphorus trihydride (PH₃)
- Phosphorus pentachloride (PCl₅)

100%

100%

Use Lewis structures to determine the molecular geometry of the following molecules. Use VSEPR to determine the molecular geometry.

Use Lewis structures to determine the molecular geometry of the following molecules. Use VSEPR to determine the molecular geometry.

Use Lewis structures to determine the molecular geometry of the following molecules. Use VSEPR to determine the molecular geometry.

- Use Lewis structures to determine the molecular geometry of the following molecules. Use VSEPR to determine the molecular geometry.

Use Lewis structures to determine the molecular geometry of the following molecules. Use VSEPR to determine the molecular geometry.

Table 1.1: Molecular Geometry of Common Compounds

Molecule	Formula	Geometry	Angle
Linear	BeH ₂	Linear	180°
Trigonal planar	BF ₃	Trigonal planar	120°
Trigonal pyramidal	BF ₃	Trigonal pyramidal	109°
Trigonal bipyramidal	PCl ₅	Trigonal bipyramidal	120°
Octahedral	SF ₆	Octahedral	90°

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Tao Wei



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