#### I. ONE REACTANT IN EXCESS - COMBUSTION

You burn an excess of butane (C<sub>4</sub>H<sub>10</sub>) in a limited amount of oxygen (O<sub>2</sub>). If 67.2 L of CO<sub>2</sub> gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:

STEP 2: Use a molar conversion to convert quantity of product given to moles:

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2$$
 \*  $(13 \text{ mol O}_2)$  =  $4.875 \text{ mol O}_2$   
 $(8 \text{ mol CO}_2)$ 

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol } O_2 \ (22.4 \text{ L} \cdot O_2) = 109.2 \text{ L} \cdot O_2 = 109 \text{ L} \cdot O_2$$
  
(1 mol  $O_2$ )

#### II. ONE REACTANT IN EXCESS - SINGLE DISPLACEMENT

14 g of lithium is combined with excess K<sub>2</sub>SO<sub>4</sub>. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP I: Balance the reaction:

Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

2 mol Li \* 
$$(1 \text{ mol Li}_2SO_4) = 1 \text{ mol Li}_2SO_4$$
  
(2 mol Li)

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

# **Mixed Stoichiometry Problems 2 And Answer**

Hsiang-Chuan Liu, Wen-Pei Sung, Wenli Yao

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