

Equilibrium HOMEWORK

Problem 1

1. Write the equilibrium constant expressions for the following reactions. Use pressures for gases and concentrations for solutions.



*Remember that SOLIDS and LIQUIDS are not included in the equilibrium expression!

2. The reaction below takes place in a 1-liter container. It produces a mixture that contains 0.45M NH_4Cl , 0.20M NH_3 , and 0.20M Cl^- . This mixture is at equilibrium at 1200 K.



a. What is the equilibrium constant for this reaction at 1200 K?

$$K = \frac{(P_{\text{NH}_3})^2 (P_{\text{Cl}_2})^2}{1} = (0.20)^2 (0.20)^2$$

$$K = 0.0016 \text{ (or } 1.6 \times 10^{-3} \text{) or } 1.6 \times 10^{-3} \text{ or } 0.0016 \text{ or } 0.0016$$

b. Based on the equilibrium constant, is this reaction reactant-favored or product-favored at 1200 K?

reactant-favored, because K is less than one, meaning there are more reactants than products at equilibrium.

3. The reaction below takes place in a 1-liter container. It produces a gaseous mixture that contains 0.35M CO , 0.40M H_2 , 0.20M H_2O , and 0.175 atm CH_4 . This mixture is at equilibrium at 1200 K.



a. What is the equilibrium constant for this reaction at 1200 K?

*Don't include liquid water in your equilibrium constant expression!
 $K = \frac{(P_{\text{CO}}) (P_{\text{H}_2})^3}{(P_{\text{CH}_4}) (P_{\text{H}_2\text{O}})}$

$$K = \frac{(0.35)(0.40)^3}{(0.175)(0.20)} = 0.32 \text{ (or } 0.32 \text{ or } 0.32 \text{ or } 0.32)$$

b. Based on the equilibrium constant, is this reaction reactant-favored or product-favored at 1200 K?

reactant-favored, because K is less than one, meaning there are more reactants than products at equilibrium.

4. The reaction below takes place in a 50.0-liter container. It produces a gaseous mixture that contains 0.00M H_2 , 0.00M H_2 , and 1.00M NH_3 . This mixture is at equilibrium at 1200 K.



a. What is the equilibrium constant for this reaction at 1200 K?

*Because the equilibrium given is as SOLID and LIQUID, given are expressed as SOLID and LIQUID, except for water.

$$K = \frac{(P_{\text{NH}_3})^2}{(P_{\text{N}_2}) (P_{\text{H}_2})^3}$$

*Remember that SOLID and LIQUID are not included in the equilibrium expression. It can be SOLIDITY, which is water divided by 55.5.
 $K = \frac{(1.00)^2}{(0.00)(0.00)^3} = 1.00 \text{ (or } 1.00 \text{ or } 1.00 \text{ or } 1.00)$

b. Based on the equilibrium constant, is this reaction reactant-favored or product-favored at 1200 K?

product-favored, because K is greater than one, meaning there are more products than reactants at equilibrium.

Mixed Review Chemical Equilibrium Answers

R Bogdan



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