

- 5) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \times 329 \times .0821}{.09 \text{ atm}} = 2310.9 \text{ L}$$

- 7) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V}$$

$$\frac{17 \cdot .0821 \cdot 340}{88.89 \text{ L}} = 5.34 \text{ atm}$$

- 8) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT}$$

$$\frac{.0821 \times 300}{.5 \text{ atm} \times 25 \text{ L}} = 1.97 \text{ mol}$$

- 9) If I have 21 moles of gas held at a pressure of 78 atm and a temperature of 900 K, what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P}$$

$$\frac{21 \cdot 0.0821 \cdot 900}{78} = 19.89 \text{ L}$$

- 10) If I have 1.9 moles of gas held at a pressure of 5 atm and in a container with a volume of 50 liters, what is the temperature of the gas?

$$PV = nRT$$

$$T = \frac{PV}{nR}$$

$$\frac{5 \text{ atm} \cdot 50 \text{ L}}{1.9 \text{ mol} \cdot .0821} = 1602.67 \text{ K}$$

- 11) If I have 2.4 moles of gas held at a temperature of 97 °C and in a container with a volume of 45 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V}$$

$$\frac{2.4 \times .0821 \cdot 370}{45} =$$

$$1.62 \text{ atm}$$

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Table of Contents Mixed Gas Laws Answers

1. Understanding the eBook Mixed Gas Laws Answers
 - The Rise of Digital Reading Mixed Gas Laws Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Gas Laws Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Gas Laws Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Mixed Gas Laws Answers
 - Personalized Recommendations
 - Mixed Gas Laws Answers User Reviews and Ratings
 - Mixed Gas Laws Answers and Bestseller Lists
5. Accessing Mixed Gas Laws Answers Free and Paid eBooks
 - Mixed Gas Laws Answers Public Domain eBooks
 - Mixed Gas Laws Answers eBook Subscription Services
 - Mixed Gas Laws Answers Budget-Friendly Options
6. Navigating Mixed Gas Laws Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Gas Laws Answers Compatibility with Devices
 - Mixed Gas Laws Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Gas Laws Answers
 - Highlighting and Note-Taking Mixed Gas Laws Answers
 - Interactive Elements Mixed Gas Laws Answers
8. Staying Engaged with Mixed Gas Laws Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Gas Laws Answers
9. Balancing eBooks and Physical Books Mixed Gas Laws Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Gas Laws Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mixed Gas Laws Answers
 - Setting Reading Goals Mixed Gas Laws Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Mixed Gas Laws Answers
 - Fact-Checking eBook Content of Mixed Gas Laws Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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