



Chemistry Chapter 10

Name _____

Stoich Worksheet 2

Stoichiometry Practice Problems 1

Do this on a separate sheet of paper. **Show all work and use units.** Round your answers to 3 significant digits.

1. What mass of iron is needed to react completely with 64.2 g of sulfur?



2. How much copper (I) oxide is needed to produce 75.0 g of copper (I) chloride?



3. How much hydrogen sulfite is produced when 178 g of sulfur dioxide reacts with water?



4. If 65.0 g of iron (II) sulfide is added to hydrogen chloride, what mass of hydrogen sulfide is produced?



5. Silver bromide is a precipitate that is formed when silver nitrate reacts with sodium bromide. What mass of the precipitate is produced if you start with 48.7 g of silver nitrate?



6. How much hydrogen nitrate is needed to react completely with 36.3 g of magnesium?



7. What mass of hydrogen can be produced when 82.1 g of zinc reacts with hydrogen sulfate?



For #8 & 9, you will need to use the density of oxygen gas is 1.43g/L.

8. What volume of oxygen is produced when 28.9g of water is decomposed?



9. What volume of oxygen is produced from the decomposition of 145.7g of potassium chlorate?



Mixed Stoichiometry Word Problems

L Darling-Hammond



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