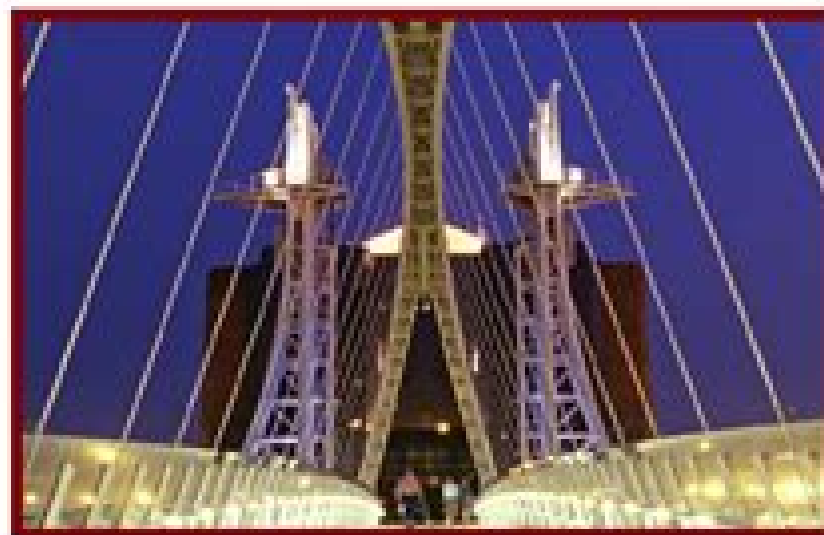


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sample answer sugar has a much greater solubility in hot water than cold water so its solubility curve probably resembles the
sloping solubility curve of potassium nitrate temperature solubility g 100 ml 10 c 36 g 100 ml 20 c 36 g 100 ml 30 c 37 g 100
ml 40 c 37 g 100 ml 50 c 38 g 100 ml 60 c 38 g 100 ml 70 c 39 g 100 ml 80

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solubility is defined as the maximum quantity of solute that can dissolve in a certain quantity of solvent at a specified
temperature differences in the solubility of different solutes are related primarily to differences in the chemical nature of the
solutes and the solvents

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1 the solubility of a solute is the maximum mass of the solute that will dissolve in a certain amount of water at a certain this
is the same as saying that solubility is the concentration of a unsaturated saturated solution of the solute 2 from your graph
what mass of kcl can be dissolved in 100 g of water at these temperatures

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