

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 kPa / 1 atm$ to get
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

or $P = 166.29 kPa$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

or $P = 540.61 kPa$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

Ideal Gas Answers Show Work

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Classical and Quantum Thermal Physics R. Prasad, 2016-11-02 Discusses the interactions of heat energy and matter

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both Conceptual clarity and problem solving skills As Maharashtra MHT CET exam emphasizes application based questions students often find it challenging to bridge the gap between theory practical problem solving This book Smash MHT CET Physics with Crash course is designed to provide a structured concise and effective approach to mastering Physics for the CET examinations This book follows strategic crash course format making it ideal for last minute Preparation as well as structured revision The content is presented in a lucid manner with emphasis on Conceptual clarity formulae shortcuts practice Problems

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Modern Engineering Thermodynamics - Textbook with Tables Booklet Robert T. Balmer, 2011-01-03 Modern Engineering Thermodynamics Textbook with Tables Booklet offers a problem solving approach to basic and applied engineering thermodynamics with historical vignettes critical thinking boxes and case studies throughout to help relate abstract concepts to actual engineering applications It also contains applications to modern engineering issues This textbook is designed for use in a standard two semester engineering thermodynamics course sequence with the goal of helping students develop engineering problem solving skills through the use of structured problem solving techniques The first half of the text contains material suitable for a basic Thermodynamics course taken by engineers from all majors The second half of the text is suitable for an Applied Thermodynamics course in mechanical engineering programs The Second Law of Thermodynamics is introduced through a basic entropy concept providing students a more intuitive understanding of this key course topic Property Values are discussed before the First Law of Thermodynamics to ensure students have a firm understanding of property data before using them Over 200 worked examples and more than 1 300 end of chapter problems provide an extensive opportunity to practice solving problems For greater instructor flexibility at exam time thermodynamic tables are provided in a separate accompanying booklet University students in mechanical chemical and general engineering taking a thermodynamics course will find this book extremely helpful Provides the reader with clear presentations of the fundamental principles of basic and applied engineering thermodynamics Helps students develop engineering problem solving skills through the use of structured problem solving techniques Introduces the Second Law of Thermodynamics through a basic entropy concept providing students a more intuitive understanding of this key course topic Covers Property Values before the First Law of Thermodynamics to ensure students have a firm understanding

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Answer Keys Detailed Solutions of all the Exercises and Past years problems are provided at the end of the chapter CBSE CLASS XI SCIENCE (PHYSICS) Study Notes | A Handbook for Class IX EduGorilla Prep Experts, *Mastering Physics for IIT-JEE Volume - I* Rathi Rakesh, Physics for IIT JEE **Thermodynamic Processes 1** Salah Belaadi,2020-02-19

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Thermodynamics of the Earth and Planets Alberto Patiño Douce,2011-08-25 This textbook provides an intuitive yet mathematically rigorous introduction to the thermodynamics and thermal physics of planetary processes It demonstrates how the workings of planetary bodies can be understood in depth by reducing them to fundamental physics and chemistry The

book is based on two courses taught by the author for many years at the University of Georgia It includes Guided Exercise boxes end of chapter problems worked solutions provided online and software boxes Maple code provided online As well as being an ideal textbook on planetary thermodynamics for advanced students in the Earth and planetary sciences it also provides an innovative and quantitative complement to more traditional courses in geological thermodynamics petrology chemical oceanography and planetary science In addition to its use as a textbook it is also of great interest to researchers looking for a one stop source of concepts and techniques that they can apply to their research problems

Thermodynamics for Chemical Engineers Kenneth Richard Hall,Gustavo Arturo Iglesias-Silva,2022-06-07

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