

7. A 645-gram beaker of water loses 12.5 kJ of water, if the water temperature of water started out at 49.5 °C, at what temperature did it end up?

a) $\Delta T = q/mc$ $\Delta T = -12,500 \text{ J} / (45 \text{ g} \cdot 4.184 \text{ J/g}^\circ\text{C}) = -4.48^\circ\text{C}$

b) $T_{\text{final}} = \Delta T + T_{\text{initial}}$ $T_{\text{final}} = -4.48^\circ\text{C} + 49.5^\circ\text{C} =$

45.0 °C

8. A 75.0 g piece of glass absorbs 6,350 joules of heat, if its initial temperature is 21.6 °C, what is its final temperature? Its specific heat capacity is 0.840 J/g °C.

a) $\Delta T = q/mc$ $\Delta T = 6,350 \text{ J} / (75.0 \text{ g} \cdot 0.840 \text{ J/g}^\circ\text{C}) = 101^\circ\text{C}$

b) $T_{\text{final}} = \Delta T + T_{\text{initial}}$ $T_{\text{final}} = 101^\circ\text{C} + 21.6^\circ\text{C} =$

123 °C

9. What is the specific heat capacity of mercury if 25.0 grams of it is heated from 25.0 °C to 155.0 °C while absorbing 455 joules of heat?

a) $\Delta T = 155.0^\circ\text{C} - 25.0^\circ\text{C} = 130.0^\circ\text{C}$

b) $c = q/m\Delta T$ $c = 455 \text{ J} / (25.0 \text{ g} \cdot 130.0^\circ\text{C}) =$

0.140 J/gg °C

10. How much heat is lost when 114 grams of lead cools from 155 °C to 22.2 °C? The specific heat capacity of lead is 0.129 J/g °C.

a) $\Delta T = 22.2^\circ\text{C} - 155^\circ\text{C} = -133^\circ\text{C}$

b) $q = (114 \text{ g})(0.129 \text{ J/g}^\circ\text{C})(-133^\circ\text{C}) =$

1,960 J or 1.96 kJ

High School Chemistry Thermochemistry Questions

Julia Schneider



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