

ANSWERS TO TOPIC 1

Review Questions

1. 1 2. 3 3. 2
4. 4 5. 2 6. 2
7. 1 8. 3 9. 3
10. The plum pudding model pictures the atom with an even distribution of positive charges, with electrons throughout the atom. The planetary model pictures the nucleus at the center of the atom with electrons in orbits.
11. Rutherford fired positively charged alpha particles at gold foil. He expected them to pass through with little deflection. When some were greatly deflected, he reasoned that they must have been repelled by a massive positively charged particle, which was called the nucleus.
12. The Rutherford model pictured the electrons in fixed, orbiting positions. The wave-mechanical model pictures electrons as possessing both wave and particle properties.
13. 3 14. 3 15. 1
16. 3 17. 4 18. 4
19. 2 20. 4 21. 2
22. 1 23. 1 24. 3
25. 3 26. 2 27. 2
28. 3 29. 3 30. 2
31. 3 32. 4 33. 4
34. 4 35. 1 36. 2
37. 1 38. 1 39. 3
40. 1 41. 2 42. 3
43. 2 44. 4 45. 4
46. 3 47. 3 48. 4
49. 2 50. 3 51. 2
52. 2 53. 3 54. 1
55. 3 56. 3
57. Na $1s^2 2s^2 2p^6 3s^1$
Mg $1s^2 2s^2 2p^6 3s^2$
Al $1s^2 2s^2 2p^6 3s^2 3p^1$
Si $1s^2 2s^2 2p^6 3s^2 3p^2$
P $1s^2 2s^2 2p^6 3s^2 3p^3$
S $1s^2 2s^2 2p^6 3s^2 3p^4$
Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Ar $1s^2 2s^2 2p^6 3s^2 3p^4$
58. 1 59. 4 60. 1
61. 1 62. 2 63. 4
64. 4 65. 2 66. 2
67. 2 68. 2 69. 1
70. The mixture contains at least three parts. One solid remained on the filter paper. The other two parts consist of a liquid with a solid dissolved in it.
71. If the calcium and sulfur can be separated by physical means, the substance is a mixture. If chemical means are required to separate them, it is a compound.
72. The picture should show several of each type of molecule evenly mixed.
73. A mixture of iron and oxygen would have the iron present as a solid and the oxygen present as a gas. The iron would be attracted to a magnet. A glowing splint would ignite into a flame if placed into the oxygen gas. If the iron and oxygen were

combined as a compound, it would be a solid that is not attracted to a magnet.

74. Pepperoni pizza is heterogeneous because individual parts can be seen.
75. A is a mixture of an element and a compound; one molecule contains identical atoms and the other one contains different types of atoms. B is a compound; all particles are identical and are made up of two different types of atoms. C is an element; all molecules are identical and each is made up of two identical atoms. D is a mixture of an element and two different compounds; one molecule contains identical atoms and the other two contain different types of atoms, but they differ from each other.
76. Container D is homogeneous because the particles are mixed on the molecular level.
77. The two substances are different materials because their composition varies.

Regents Practice Questions

Part A

- | | | |
|-------|-------|-------|
| 1. 1 | 2. 1 | 3. 1 |
| 4. 3 | 5. 1 | 6. 4 |
| 7. 2 | 8. 3 | 9. 3 |
| 10. 2 | 11. 3 | 12. 1 |
| 13. 2 | 14. 3 | |

Part B

- | | | |
|-------|-------|-------|
| 15. 4 | 16. 3 | 17. 1 |
| 18. 3 | 19. 1 | 20. 3 |
| 21. 1 | 22. 3 | 23. 2 |
| 24. 4 | | |

Part C

25. Because the two nuclei have different numbers of protons, they have different amounts of attraction for electrons. As a result, the amounts of energy separating the energy levels differ for the two elements.
26. The samples are not the same substance because the percent of each element differs from sample to sample.
27. Because he assumed the atom to have an even distribution of positive charges, he did not expect a concentration of positive charges that would strongly deflect or bounce back the alpha particles.
28. The spectra are based on the elements present in the compound. If some lines matched and some didn't, the compounds have a least one element in common and at least one differing element.
29. Column 1: 1, 2, 1; column 2: K, 39, 19; column 3: 92, 146, 92
30. $(0.9778)(12) + (0.0222)(13) = 12.02$
31. They contain different numbers of neutrons.
32. $(19.99 \times .909) + (20.99 \times 0.03) + (21.99 \times 0.88)$
33. 20
34. $\frac{(12.00 \times 9.93) + (13.00 \times 1.07)}{100}$
35. Both isotopes contain 6 protons but different numbers of neutrons. Or: An isotope with a mass of 12.00 contains 6 protons and 6 neutrons, while the isotope with a mass of 13.00 contains 6 protons and 7 neutrons.

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Chapter 2 Answer Key

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Table of Contents High Marks Regents Living Environment Made Easy Chapter 2 Answer Key

1. Understanding the eBook High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - The Rise of Digital Reading High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Personalized Recommendations
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key User Reviews and Ratings
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key and Bestseller Lists
5. Accessing High Marks Regents Living Environment Made Easy Chapter 2 Answer Key Free and Paid eBooks
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key Public Domain eBooks
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key eBook Subscription Services
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key Budget-Friendly Options
6. Navigating High Marks Regents Living Environment Made Easy Chapter 2 Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key Compatibility with Devices
 - High Marks Regents Living Environment Made Easy Chapter 2 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Highlighting and Note-Taking High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Interactive Elements High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
8. Staying Engaged with High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
9. Balancing eBooks and Physical Books High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
- Setting Reading Goals High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
- Fact-Checking eBook Content of High Marks Regents Living Environment Made Easy Chapter 2 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements

- Interactive and Gamified eBooks

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