

Chapter 2

The First Law of Thermodynamics

2.2. Work is defined as *negative* $p\Delta V$ because if a system does work on the surroundings, the system *loses* energy.

$$2.4. \quad 1780 \text{ torr} \left| \frac{1 \text{ atm}}{(760 \text{ torr})} \right| = 2.34 \text{ atm}$$

$$w = -P\Delta V = -2.34 \text{ atm} (1.00 \text{ L} - 3.55 \text{ L}) \left(\frac{101.32 \text{ J}}{1 \text{ L} \cdot \text{atm}} \right) = +605 \text{ J}$$

605 J of work are done ON the system.

- 2.6. (a) The work would be less because the external pressure is less.
(b) The work would be greater because the external pressure is greater.
(c) No work would be performed because the external pressure is (effectively) zero.

2.8. First, we need to find the final volume of the CO_2 . Using $P_1V_1 = P_2V_2$, we find that $V_2 = 105 \text{ mL}$.

$$w = -(1.0 \text{ atm})(105 - 25.0 \text{ mL}) \left(\frac{1 \text{ L}}{1000 \text{ mL}} \right) \left(\frac{101.32 \text{ J}}{1 \text{ L} \cdot \text{atm}} \right) = -8.1 \text{ J}$$

2.10. First, determine the ΔT : $330 \text{ K} - 298 \text{ K} = 32 \text{ K}$.

$$\text{Using } q = m \cdot c \cdot \Delta T, \text{ we rearrange: } c = \frac{q}{m \cdot \Delta T}$$

$$\text{Substituting: } c = \frac{288 \text{ J}}{(50.5 \text{ g})(32 \text{ K})} = 0.178 \frac{\text{J}}{\text{g} \cdot \text{K}}$$

2.12. First, we need to calculate the number of moles of phosphorous:

$$\left(\quad \right)$$

Gases Solution Manual

**Katherine J. Denniston, Denniston
Katherine**



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INTRODUCTION

1 1 1 Purpose This chapter provides a general history of the development of military diving operations

1 1 2 Scope This chapter outlines the hard work and dedication of a number of individuals who were pioneers in the development of diving technology As with any endeavor it is important to build on the discoveries of our predecessors and not repeat mistakes of the past

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