

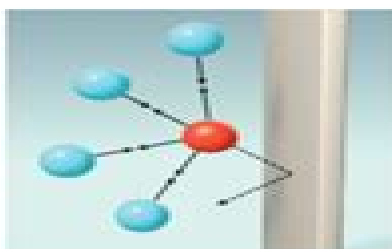
Chapter 11: Gases

Focus of this section:

Deviation from ideality

Deviation from Ideal Behavior:

- recall: real gases deviate from ideal gas behavior at high pressure and low temperature conditions
 1. real molecules have volume
 - cannot compress gas to zero volume
 2. real molecules interact (attracted to each other)
 - causes lower P (vs. ideal gas), especially at lower temperatures



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$$= P_{\text{obs}} < P_{\text{ideal}}$$

= need to add correction to pressure term

= Since 2 molecules cannot occupy the same space, $V_{\text{accessible}} < V_{\text{container}}$ and a corrective term should be subtracted from V in ideal gas equation

$$\left(\underbrace{P + \frac{an^2}{V^2}}_{\text{corrected pressure term}} \right) \left(\underbrace{V - nb}_{\text{corrected volume term}} \right) = nRT$$

$PV = nRT$

$$P_{\text{ideal}} = \left(P_{\text{obs}} + \frac{n^2 a}{V^2} \right)$$

$$\text{correction} \propto \left(\frac{n}{V} \right)^2$$

$$V_{\text{ideal}} = (V_{\text{container}} - nb)$$

$$\text{correction} \propto n$$

= Note, "a" and "b" are different for each gas

Gases Section Quizzes

Rachel Sandford



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