

9. A 12.0 L sample of nitrogen has a pressure of 540 kPa at 150 °C. The gas is compressed to 1.20 L and the temperature is lowered to 50 °C. What is the new pressure?

$$\frac{540 \text{ kPa} \cdot 12.0 \cancel{\text{L}}}{423 \text{ K}} = \frac{P_2 \cdot 1.20 \cancel{\text{L}}}{323 \text{ K}}$$

$$P_2 = 4,123 \text{ kPa}$$

10. What mass of helium is needed to inflate a balloon to 4.0 L tank at 305 K and 2.5 atm?

$$n = \frac{PV}{RT} = \frac{(2.5 \text{ atm})(4.0 \cancel{\text{L}})}{(0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}})(305 \text{ K})} = 0.399 \text{ mol He} \left| \frac{4 \text{ g He}}{1 \text{ mol He}} \right| = 1.6 \text{ g He}$$

11. Compare the diffusion rates of helium and radon.

12. A sample of ozone is collected over water at 27 °C and at a barometric pressure of 765 torr. What is the partial pressure of the ozone gas?

$$P_{\text{O}_2} = 765 \text{ torr} - 26 \text{ torr} = 739 \text{ torr}$$

13. Are volume and temperature inversely or directly related? Give a real life example.

$\frac{V}{T}$ As Temp. inc., vol. inc.
when pouring hot water over balloon, balloon expands

14. A mixture of gases contains N₂, O₂, and CO₂. If the total pressure of the mixture is 10.0 atm and N₂ = 2.0 atm and O₂ = 3.5 atm, what is the partial pressure of CO₂?

$$10.0 \text{ atm} = 2.0 \text{ atm} + 3.5 \text{ atm} + P_{\text{O}_2}$$

$$(4.5 \text{ atm} = P_{\text{O}_2})$$

15. If a sample contains 8 moles of H₂, 3 moles of O₂, and 2 moles of He in a 7 liter vessel at 19 °C, determine the partial pressure of each gas and the total pressure of the mixture in units of atm.

$$P_{\text{H}_2} = \frac{(8 \text{ mol H}_2)(0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}})(292 \text{ K})}{7 \text{ L}} = 27.4 \text{ atm}$$

$$P_{\text{O}_2} = \frac{(3 \text{ mol O}_2)(0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}})(292 \text{ K})}{7 \text{ L}} = 10.3 \text{ atm}$$

$$P_{\text{He}} = \frac{(2 \text{ mol He})(0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}})(292 \text{ K})}{7 \text{ L}} = 6.84 \text{ atm}$$

$$P_{\text{T}} = 44.5 \text{ atm}$$

16. If a gas sample with a pressure of 86.2 kPa contains 53% methane gas, what is the partial pressure of methane?

$$86.2 \text{ kPa} = 45.7 \text{ kPa} + 40.5 \text{ other gas}$$

$$P_{\text{CH}_4} = (0.53)(86.2 \text{ kPa}) = 45.7 \text{ kPa}$$

17. If 2.3 L of hydrogen reacts with oxygen to form water at STP, how many grams of oxygen are reacting?



$$2.3 \cancel{\text{L}} \text{ H}_2 \rightarrow 1.15 \text{ L O}_2 \text{ reacts}$$

$$\frac{2.3 \cancel{\text{L}} \text{ H}_2}{2 \cancel{\text{L}} \text{ H}_2} \times \frac{1 \text{ L O}_2}{1 \cancel{\text{L}} \text{ H}_2} = 1.15 \text{ L O}_2$$

$$1.05 \text{ g O}_2$$

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SA Adler

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Table of Contents Gases Chem Answer Key

1. Understanding the eBook Gases Chem Answer Key
 - The Rise of Digital Reading Gases Chem Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Chem Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Chem Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Chem Answer Key

- Personalized Recommendations
 - Gases Chem Answer Key User Reviews and Ratings
 - Gases Chem Answer Key and Bestseller Lists
5. Accessing Gases Chem Answer Key Free and Paid eBooks
 - Gases Chem Answer Key Public Domain eBooks
 - Gases Chem Answer Key eBook Subscription Services
 - Gases Chem Answer Key Budget-Friendly Options
 6. Navigating Gases Chem Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases Chem Answer Key Compatibility with Devices
 - Gases Chem Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases Chem Answer Key
 - Highlighting and Note-Taking Gases Chem Answer Key
 - Interactive Elements Gases Chem Answer Key
 8. Staying Engaged with Gases Chem Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Chem Answer Key
 9. Balancing eBooks and Physical Books Gases Chem Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Chem Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gases Chem Answer Key
 - Setting Reading Goals Gases Chem Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gases Chem Answer Key

- Fact-Checking eBook Content of Gases Chem Answer Key
 - Distinguishing Credible Sources
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- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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