

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$$PV = nRT$$

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$$

$$2.34 \text{ L Cl}_2$$



- a. Balance the above reaction.

- b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$$n = x$$

$$V = 0.628 \text{ L}$$

$$T_1 = 25 + 273 = 298 \text{ K}$$

$$P_1 = 1.05 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.628 \text{ mol}) (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (298 \text{ K})}{1.05 \text{ atm}} = 14.6 \text{ L}$$

3. Potassium permanganate,  $\text{KMnO}_4$ , is produced commercially by oxidizing aqueous potassium manganate,  $\text{K}_2\text{MnO}_4$ .



- a. Balance the above reaction.

- b. What volume of  $\text{Cl}_2(g)$ , measured at STP, is needed to produce 10.0 g of  $\text{KMnO}_4$ ?

$$V = \frac{nRT}{P}$$

$$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mole}}{158 \text{ g}} \right| \left| \frac{1 \text{ mole Cl}_2}{2 \text{ moles KMnO}_4} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mole Cl}_2} \right| = 0.709 \text{ L Cl}_2$$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.

- b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$$PV = nRT$$

$$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mole}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{1 \text{ mole}} \right| = 0.204 \text{ mol}$$

$$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mole}}{18 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{3 \text{ moles H}_2\text{O}} \right| = 0.381 \text{ mol}$$

$$T_1 = 24 + 273 = 297 \text{ K}$$

$$P_1 = 752 / 760 = 0.989 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.204 \text{ mol}) (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (297 \text{ K})}{0.989 \text{ atm}} = 5.03 \text{ L}$$

# Gas Stoichiometry Answer Key

**Xiaolong Qi**



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