

Name \_\_\_\_\_

Section \_\_\_\_\_

Date \_\_\_\_\_

20. A gas sample occupies 200 mL at 700 mm Hg. What volume does the gas occupy at 600 mm Hg?

10.00

$$(200 \text{ mm Hg})(200 \text{ mL}) = (600 \text{ mm Hg}) V_2$$

$$V_2 = 300 \text{ mL}$$

21. Nitrogen gas in a sealed container exerts a pressure of 150 atm at 300 K. What will the pressure be if the temp. for 100 atm is 323 K in the same seal that increased temperature from 300 K?

10.00

$$\frac{150 \text{ atm}}{300 \text{ K}} = \frac{P_2}{323 \text{ K}}$$

$$P_2 = 164 \text{ atm}$$

22. A sample of gas occupies a volume of 80 mL at a pressure of 0.50 atm and a temperature of 273 K. What will be the volume at a pressure of 1.50 atm and a temperature of 323 K?

10.00

$$\frac{(0.50 \text{ atm})(80 \text{ mL})}{273 \text{ K}} = \frac{(1.50 \text{ atm}) V_2}{323 \text{ K}}$$

$$V_2 = 31.55 \text{ mL}$$

# Gas Laws Test Review Answers

**Kaplan**



## Gas Laws Test Review Answers:

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