

Chapter 14—WS GAS LAWS

1. A quantity of gas under a pressure of 800 mm Hg has a volume of 380 liters. What is its volume at standard pressure when temperature is held constant?
2. A quantity of gas has a volume of 120 liters when confined under a pressure of 700.0 mm Hg at a temperature of 20.0°C. At what pressure will its volume be 30.0 liters at 20.0°C?
3. At constant pressure, the volume of a gas is increased from 150 liters to 300.0 liters by heating it. If the original temperature of the gas was 20.0°C, what will its final temperature be?
4. A quantity of gas exerts a pressure of 740 mm Hg at a temperature of 22.0°C. If the volume remains unchanged, what pressure will it exert at -8.0°C?
5. A quantity of oxygen gas has a volume of 850 ml when measured at 27.0°C and 730 mm Hg of pressure. Determine its volume at STP.
6. When measured at STP, a quantity of gas has a volume of 500.0 liters. What volume will it occupy at 0.0°C and a pressure of 700.0 mm Hg?
7. Know the relationships between your four variables- pressure, volume, temperature, and moles. Which are inverse and which are direct?? Be able to explain them in terms of what is happening to the energy, collisions, and space between the particles.
8. A gas whose behavior closely resembles that of an ideal gas has a volume of 2.00 liters at a temperature of 27.0°C and a pressure of 900.0 mm Hg. (a) How many moles of molecules are in the sample? (b) How many molecules are in the sample?
9. Know Graham's law of effusion. Be able to calculate the ratio of rate of effusion. Example: what is the rate of effusion of oxygen gas compared to nitrogen if both gases are at the same temperature?

10. How many moles?

$$1. P_1 V_1 = P_2 V_2$$

$$800 \text{ mmHg} (380 \text{ L}) = 760 \text{ mmHg} (V_2)$$

$$V_2 = 400 \text{ L}$$

$P \downarrow V \uparrow$ checks out!

$$2. P_1 V_1 = P_2 V_2$$

$$700.0 \text{ mmHg} (120 \text{ L}) = P_2 (30 \text{ L})$$

$$P_2 = 2800 \text{ mmHg}$$

$V \downarrow P \uparrow$ checks out!

$$3. \frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \frac{150 \text{ L}}{293 \text{ K}} = \frac{300 \text{ L}}{T_2}$$

$$T_2 (150) = 300 (293)$$

$$T_2 = 586 \text{ K}$$

$V \uparrow T \uparrow$
checks out!

Gas Laws Review Sheet Answers

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