

Worksheet – Gas Laws III

Gas Laws: Combined Gas Law

The following problems may be solved using the relationships described by either Charles' or Boyles' laws, together in the combined gas law. Show all the steps to the solution of the problem.

$$V_1P_1/T_1 = V_2P_2/T_2$$

1. Carbon dioxide occupies a 2.54 dm³ container at STP. What will be the volume when the pressure is 150 KPa and 26°C?

$$V_2 = \frac{[2.54 \text{ dm}^3][101.3 \text{ KPa}][26^\circ\text{C} + 273\text{K}]}{[150 \text{ KPa}][273 \text{ K}]} = 1.88 \text{ dm}^3$$

2. Oxygen occupies a fixed container of 5.5L at STP. What will happen to the pressure if the temperature rises to 300K?

$$P_2 = \frac{[5.5 \text{ L}][101.3 \text{ KPa}][300 \text{ K}]}{[5.5 \text{ L}][273 \text{ K}]} = 111.3 \text{ KPa}$$

3. Methane is compressed in a closed 15.8 dm³ container at 101.3 KPa. If the volume drops to 8.7 dm³ and the temperature begins at 25°C and then drops to 18°C , what will the pressure of the gas be?

$$P_2 = \frac{[101.3 \text{ KPa}][15.8 \text{ dm}^3][300 \text{ K}]}{[8.7 \text{ dm}^3][273 \text{ K}]} = 179.6 \text{ KPa}$$

4. A helium balloon is fully inflated at 1.2 L. When the clerk is filling the balloon, she stops to make sure it is not going to explode and checks the pressure. The pressure is 5.7 KPa and the temperature at the store is 24°C. The balloon is 0.75L full. She stops when the balloon is fully inflated and the pressure is 7.2 KPa. When the customer takes the balloon outside, it explodes. What was the temperature outdoors?

$$T = \frac{[24^\circ\text{C} + 273\text{K}][1.2 \text{ L}][7.2 \text{ KPa}]}{[0.75 \text{ L}][5.7 \text{ KPa}]} = 600 \text{ K}$$

5. Oxygen gas is added to a rebreather at the same rate as it would be in the open air. The volume is 4.8 dm³ and the temperature is 25 °C. The pressure is 1 atm. Once the tank is lowered under the water the pressure increases to 2 atm and the temperature drops to 12°C. Is the tank in danger of exploding if it's maximum volume is 5.0 dm³?

$$V_2 = \frac{[4.8 \text{ dm}^3][1 \text{ atm}][12^\circ\text{C} + 273\text{K}]}{[2 \text{ atm}][25^\circ\text{C} + 273\text{K}]} = 2.1 \text{ dm}^3, \text{ not in danger of exploding}$$

6. Ozone is formed due to electricity passing through the air and splitting some O₂ atoms which then join briefly to form O₃, ozone. Your olfactory sense can detect 1.00 mL ozone in 10 L of air at STP. You are surprised that the noreaster is accompanied by

Gas Laws And Scuba Diving Worksheet

Christian Drosten

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Gas Laws And Scuba Diving Worksheet

1. Understanding the eBook Gas Laws And Scuba Diving Worksheet
 - The Rise of Digital Reading Gas Laws And Scuba Diving Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws And Scuba Diving Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Gas Laws And Scuba Diving Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws And Scuba Diving Worksheet
 - Personalized Recommendations
 - Gas Laws And Scuba Diving Worksheet User Reviews and Ratings

- Gas Laws And Scuba Diving Worksheet and Bestseller Lists
- 5. Accessing Gas Laws And Scuba Diving Worksheet Free and Paid eBooks
 - Gas Laws And Scuba Diving Worksheet Public Domain eBooks
 - Gas Laws And Scuba Diving Worksheet eBook Subscription Services
 - Gas Laws And Scuba Diving Worksheet Budget-Friendly Options
- 6. Navigating Gas Laws And Scuba Diving Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws And Scuba Diving Worksheet Compatibility with Devices
 - Gas Laws And Scuba Diving Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws And Scuba Diving Worksheet
 - Highlighting and Note-Taking Gas Laws And Scuba Diving Worksheet
 - Interactive Elements Gas Laws And Scuba Diving Worksheet
- 8. Staying Engaged with Gas Laws And Scuba Diving Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws And Scuba Diving Worksheet
- 9. Balancing eBooks and Physical Books Gas Laws And Scuba Diving Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws And Scuba Diving Worksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws And Scuba Diving Worksheet
 - Setting Reading Goals Gas Laws And Scuba Diving Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws And Scuba Diving Worksheet
 - Fact-Checking eBook Content of Gas Laws And Scuba Diving Worksheet
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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