

# Gas Laws Problem Packet

**Gas Laws Packet**

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

**Instructions**

Standard Conversions

Unit	Standard Pressure
atm	101.3
mmHg	760
gms	1
at	14.7

**22.4 L = 1 mol at STP**

STP (standard temperature and pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law	$P_1V_1 = P_2V_2$
Charles's Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Gramm's Law	$\frac{P_1}{P_2} = \frac{n_1}{n_2}$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- if kPa,  $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$
- if atm,  $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

**Notes:**

- Temperature must always be in Kelvin!  $K = ^\circ\text{C} + 273$
- If you are given two different units of pressure, CONVERT (one).
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you are given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (Mol/L or mol/L)
- Gramm's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

**Charles' Law**  $V_1/T_1 = V_2/T_2$

Step 1: Compare each of the problems below. How many variables are given in each problem?

- The temperature under my refrigerator is about 40°C. The gas in the tank has the temperature of 40°C and is under 100 psi. What volume of gas is in the tank?
- A balloon of gas is at 100 psi and 40°C. What volume of gas is in the tank?

**Gay-Lussac's Law**  $P_1/T_1 = P_2/T_2$

Step 1: Compare each of the problems below. How many variables are given in each problem?

- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?
- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?

**Boyle's Law**  $P_1V_1 = P_2V_2$

Step 1: Compare each of the problems below. How many variables are given in each problem?

- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?
- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?

**Avogadro's Law**  $V_1/n_1 = V_2/n_2$

Step 1: Compare each of the problems below. How many variables are given in each problem?

- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?
- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?

**Combined Gas Law**  $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

Step 1: Compare each of the problems below. How many variables are given in each problem?

- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?
- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?

**Gramm's Law**  $\frac{P_1}{P_2} = \frac{n_1}{n_2}$

Step 1: Compare each of the problems below. How many variables are given in each problem?

- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?
- A gas in a tank is at 100 psi and 40°C. What volume of gas is in the tank?

All Gas Laws Included!

# Gas Law Problems Packet

**CO Houle**



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