

UNIT 7: Gas Laws Review Sheet

1. Equations to Know:

- Dalton's Law:** _____
Total Pressure = Sum of the partial pressures and each pressure exerts a separate pressure independent of the other gases
- Boyle's Law:** _____
Volume and pressure of a gas are _____ proportional
- Charles' Law:** _____
Volume and temperature of a gas are _____ proportional
- Gay-Lussac's Law:** _____
Temperature and pressure of a gas are _____ proportional
- Ideal Gas Law:** _____
 $P = \text{Pressure, } V = \text{volume, } n = \text{number of moles, } R = \text{Universal Gas Constant, } T = \text{Temp in Kelvin}$

2. Definitions to Know

- Three Common States of Matter in Terms of Particle Arrangement, Motion, Energy and Forces
- Phase Changes (i.e., melting, freezing, etc.)
- Pressure (also know the different units for pressure)
- Heat
- Temperature (also know the two units discussed in class)
- Absolute Zero

3. What does STP stand for and what are its numerical values?

4. What are the three assumptions of the Kinetic Theory?

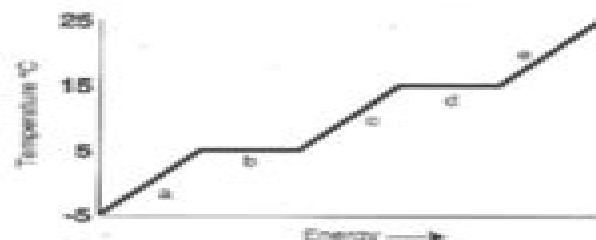
5. What are the assumptions of Ideal Gases?

6. Be able to label the phases/phase changes on a diagram similar to the one below:



7. Know the boiling point and freezing point of water.

8. If given a phase diagram like the one below, be able to identify when melting and freezing occur and answer the following the questions.



- What is the melting temperature of the substance?
- What is the boiling temperature of the substance?
- What is happening to the temperature at a, c and e?
- What is happening at points b and d?
- Label solid, liquid and gas on the curve.
- On which point of the curve are the particles moving the least?
- Where on the curve is ONLY a liquid present?

Gas Laws Unit 9 Chemistry Review Key

A Gutmann



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