

Name _____

Gas Laws

21. Q: What is the relationship between volume and moles in Avogadro's Principle?

A: Volume is directly proportional to the number of moles.

22. Q: Calculate the pressure of 1.0 mole of gas in a 10.0 L container at 273 K using the ideal gas law.

A: The pressure is 2.24 atm.

23. Q: According to the ideal gas law, what happens to the pressure if the volume is reduced by half while the temperature and number of moles remain constant?

A: The pressure doubles.

24. Q: If the volume of a gas is tripled and the temperature is doubled, what happens to the pressure according to the combined gas law?

A: The pressure decreases to 2/3 of its original value.

25. Q: Using gas stoichiometry, how many liters of CO_2 are produced from 5.0 L of O_2 in the reaction $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ at STP?

A: 5.0 L of CO_2 are produced.

26. Q: In the reaction $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$, how many liters of NO_2 are produced from 4.0 L of NO at STP?

A: 4.0 L of NO_2 are produced.

27. Q: Calculate the volume of N_2 gas required to produce 20.0 L of NH_3 in the reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ at STP.

A: 10.0 L of N_2 are required.

28. Q: If the temperature of a gas is increased from 300 K to 600 K and the volume is kept constant, what happens to the pressure?

A: The pressure doubles.

29. Q: Using the ideal gas law, how do you calculate the volume of gas if you know the pressure, temperature, and number of moles?

A: $V = nRT/P$

30. Q: In the reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$, how many liters of HCl are produced from 3.0 L of H_2 at STP?

A: 6.0 L of HCl are produced.



Gas Laws Honors Study Guide

**Elliott M. Epstein, Jerome
Shostak, Lawrence M. Troy**



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science students in order to solve quantitative word problems is important for two reasons First it challenges the practice by teachers that one specific method should be used when solving problems Secondly it calls into question the belief that multiple methods would confuse students and retard their learning Using a case study approach and informed by research conducted by the author this book claims that providing students with a choice of methods as well as requiring additional methods as a way to validate results can be beneficial to student learning A close reading of the literature reveals that time spent on elucidating concepts rather than on algorithmic methodologies is a critical issue when trying to have students solve problems with understanding It is argued that conceptual understanding can be enhanced through the use of multiple methods in an environment where students can compare evaluate and verbally discuss competing methodologies through the facilitation of the instructor This book focuses on two very useful methods proportional reasoning PR and dimensional analysis DA These two methods are important because they can be used to solve a large number of problems in all of the four academic sciences biology chemistry physics and earth science This book concludes with a plan to integrate DA and PR into the academic science curriculum starting in late elementary school through to the introductory college level A challenge is presented to teachers as well as to textbook writers who rely on the single method paradigm to consider an alternative way to teach scientific problem solving

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