| Varme | | | | |
|--------|------|---|---|--|
| Period | Date | 1 | / | |

12 • The Gas Laws

THE COMBINED GAS LAW

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$
 or $P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$
 $K = {}^{\circ}C + 273$

Serater

Complete the following chart.

| | P ₁ | $\mathbf{V_1}$ | T ₁ | P ₂ | V_2 | T ₂ |
|---|----------------|----------------|-------------------|----------------|---------|----------------------|
| 1 | 1.50 atm | 3.00 L | 293K | 2.50 atm | | 3051€ -30.0°C |
| 2 | 720. torr | 256. mL | 298 K -25.0°€- | | 250. mL | 3 23 K -50.0 ℃ |
| 3 | 600. mmHg | 2.50 L | 325 K -22.0°€ | 760. mmHg | 1.80 L | |
| 4 | | 750. mL | 273 K | 2.00 atm | 500. mL | 248 K -25.0°C |
| 5 | 95.0 kPa | 4.00 L | | 101. kPa | 6.00 L | 471. K or 498. °C |
| 6 | 650. torr | 8 | 373 K -100.℃ | 900. torr | 225. mL | 423 K -150. °€ |
| 7 | 850. mmHg | 1.50 L | 288 K -15.0°€ | | 2.50 L | 303 K |
| 8 | 125, kPa | 125. mL | | 100. kPa | 100 mL | 348 K |

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