

INTERMOLECULAR FORCES – ANSWER KEY

(from *Tro, Chapter 11, page 512*)

49. Determine the kinds of intermolecular forces that are present in each of the following elements or compounds:

- | | | | |
|----------------------|--------------------------|------------------------|------------------------|
| (a) Kr* | (b) NCl ₃ ** | (c) SiH ₄ * | (d) HF**** |
| (e) N ₂ * | (f) NH ₃ **** | (g) CO** | (h) CCl ₄ * |

50. Determine the kinds of intermolecular forces that are present in each of the following elements or compounds:

- | | | | |
|------------------------|--------------------------|----------------------------|----------------------|
| (a) HCl** | (b) H ₂ O**** | (c) Br ₂ * | (d) He* |
| (e) PH ₃ ** | (f) HBr** | (g) CH ₃ OH**** | (h) I ₂ * |

* dispersion forces only; ** dispersion and dipole-dipole forces; **** dispersion force, dipole-dipole forces, and hydrogen bonds

51. Arrange the following in order of increasing boiling point. Explain your reasoning.

- | | | | |
|---------------------|-------------------------------------|----------------------------------------|-------------------------------------------------------|
| (a) CH ₄ | (b) CH ₃ CH ₃ | (c) CH ₃ CH ₂ Cl | (d) CH ₃ CH ₂ OH a b c d |
|---------------------|-------------------------------------|----------------------------------------|-------------------------------------------------------|

52. Arrange the following in order of increasing boiling point. Explain your reasoning.

- | | | |
|----------------------|-----------------------|-----------------------------------|
| (a) H ₂ S | (b) H ₂ Se | (c) H ₂ O a b c |
|----------------------|-----------------------|-----------------------------------|

53. For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning.

- | | |
|------------------------------------------------------------|-------------------------------------------------------------------------------|
| (a) CH₃OH or CH ₃ SH | (b) CH ₃ OCH ₃ or CH₃CH₂OH |
| (c) CH ₄ or CH₃CH₃ | |

54. For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning.

- | | | |
|----------------------------------------------|----------------------------------------------|----------------------------------------------|
| (a) NH₃ or CH ₄ | (b) CS₂ or CO ₂ | (c) CO ₂ or NO₂ |
|----------------------------------------------|----------------------------------------------|----------------------------------------------|

55. For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning.

- | | | |
|---------------------------------------------|-----------------------------------------------|----------------------------------------------|
| (a) Br₂ or I ₂ | (b) H₂S or H ₂ O | (c) NH ₃ or PH₃ |
|---------------------------------------------|-----------------------------------------------|----------------------------------------------|

56. For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning.

- (a) **CH₄** or CH₃Cl
- (b) CH₃CH₂CH₂OH or **CH₃OH**
- (c) CH₃OH or **H₂CO**

57. Which of the following pairs of substances would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved.

- | | |
|-------------------------------------------------------------------|------------------------------------------------------------------------|
| (a) CCl ₄ and H ₂ O | (b) KCl and H₂O (ion-dipole forces) |
| (c) Br₂ and CCl₄ (dispersion forces) | (d) CH₃CH₂OH and H₂O (H-bonds) |

58. Which of the following pairs of compounds would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved.

- (a) **CH₃CH₂CH₂CH₂CH₃ and CH₃CH₂CH₂CH₂CH₂CH₃ (dispersion forces)**
- (b) CBr₄ and H₂O
- (c) **LiNO₃ and H₂O (ion-dipole forces)**
- (d) CH₃OH and CH₃CH₂CH₂CH₂CH₃

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