

# Stoichiometry with a Twist

- Given the density of oxygen is 1.439 grams per liter, how many liters of oxygen gas can be produced if 15.0g of mercury (II) oxide are heated to produce mercury and oxygen gas?



$$\begin{array}{c}
 ? \text{ L O}_2 = 15.0\text{g HgO} \left| \begin{array}{c} 1 \text{ mol HgO} \\ 217\text{g HgO} \end{array} \right| \begin{array}{c} 1 \text{ mol O}_2 \\ 2 \text{ mol HgO} \end{array} \left| \begin{array}{c} 32\text{g O}_2 \\ 1 \text{ mol O}_2 \end{array} \right| \begin{array}{c} 1 \text{ L O}_2 \\ 1.439\text{g O}_2 \end{array}
 \end{array}$$

$$= 0.770 \text{ L O}_2$$

# Episode 803 Stoichiometry

**Mike Jess**



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