

A crystalline solid is formed when 2 or more atoms exchange valence electrons to increase stability.

Ionic bond is formed when valence electrons are transferred from one atom to another.

cations = atoms lose electrons to become [+] charged ions.

atoms gain electrons to become [-] charged.

In ionic compounds that ions are arranged in a crystal lattice. Strong forces hold the ions together.

Properties of ionic compounds:

- = high melting and boiling points
- = hard & not easily crushed
- = conduct electricity when melted or dissolved because the ions are free to move

covalent bond = electrons are shared. Forming molecules

Covalent compounds have weak forces holding the molecules together.

Properties of covalent compounds:

- = low melting and boiling points
- = Many covalent compounds are volatile liquids or gases.
- = soft, easier to crush
- = are not conductors of electricity
- = electronegativity = property that tells how strong an

# Episode 5intro To Bonding

**Michael Rajnik**



## **Episode 5intro To Bonding:**

## Unveiling the Power of Verbal Artistry: An Emotional Sojourn through **Episode 5intro To Bonding**

In a world inundated with monitors and the cacophony of quick conversation, the profound energy and emotional resonance of verbal art often disappear in to obscurity, eclipsed by the constant assault of sound and distractions. Yet, nestled within the lyrical pages of **Episode 5intro To Bonding**, a captivating work of fictional splendor that pulses with natural thoughts, lies an wonderful trip waiting to be embarked upon. Written with a virtuoso wordsmith, this exciting opus courses readers on a psychological odyssey, lightly revealing the latent potential and profound influence stuck within the complicated web of language. Within the heart-wrenching expanse with this evocative evaluation, we can embark upon an introspective exploration of the book is main styles, dissect its charming writing style, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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