

Chemistry EOC Review – Answer Key

- 1) a. 68,200,000 cg b. 5000 mm c. 0.0125 kL d. 0.548 dm
- 2) ~21.5 mL
- 3) production of heat & light, formation of a gas, formation of a precipitate, color change
- 4) $33.64 \text{ g} / 4.28 \text{ ml} = 7.86 \text{ g/ml} = \text{Fe}$
- 5) m.p. $-183^{\circ}\text{C} = \text{CH}_4$ (methane)
- 6) H_2 – splint test (burning splint ignites hydrogen)
 O_2 – splint test (oxygen reignites glowing splint)
 H_2O – cobalt chloride paper (paper changes from blue to pink in the presence of water vapor)
 CO_2 – splint test (carbon dioxide extinguishes burning splint);
limewater test (limewater turns cloudy when carbon dioxide is bubbled through it,
producing CaCO_3)
- 7) The protons are positively charged particles found inside the nucleus along with the neutrons, which have no charge. The total number of these nucleons in an atom is the mass # / atomic mass of the atom. The electrons are the negatively charged particles found orbiting in “clouds” around the nucleus. Electrons can be gained by the atom to form anions or lost to form cations.
- 8) The thing that makes any element on the periodic table different from all other elements is the number of protons in its nucleus. Isotopes are atoms of the same type of element which have different masses because they have different numbers of neutrons.
- 9) a. $^{39}_{19}\text{K}^+$
b. $^{57}_{26}\text{Fe}$
c. $^{33}_{16}\text{S}^{2-}$
- 10)

<u>chemical symbol</u>	<u>atomic number</u>	<u>mass number</u>	<u># of protons</u>	<u># of electrons</u>	<u># of neutrons</u>	<u>net charge</u>
K^+	19	42	19	18	23	+1
Sr	38	88	38	38	50	0
F^-	9	20	9	10	11	-1
P	15	31	15	15	16	0
Ca^{2+}	20	42	20	18	22	+2

- 11) uranium – 238 $238 - 92 = 146$ neutrons
- 12) (1) 199.59g
(2) They contain the same # of protons (they are the same element), but different #s of neutrons (electrons don't matter, b/c they have virtually no mass), & thus different masses.

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