

## **SEMESTER 2 STUDY GUIDE**

**\*\*This review guide is intended to provide you with general topics to look over and study. Using these topics you can then go further in depth using your notes, labs, and homework assignments to study. Remember you have a book; I suggest using it. Breathe, study a bit every day, and you'll do fine\*\***

Materials Provided for Exam:

- P-Table
- Charge Card
- Mole/Stoich Road Map
- Metal Reactivity list

You must bring a Calculator- I HAVE NONE TO GIVE YOU!!!

Graphing/Data: Be able to interpret graphs and/or graph data.

### **Bonding- Ch 7&8** **Pages 204-272**

What is the difference between the following types of bonds? Give an example.  
single covalent

double covalent

triple covalent

ionic

polar covalent

non-polar covalent

What is a monatomic gas? Diatomic? Why do they form this way?

Why do hydrogen and any halogens form diatomic gases? Hint: see electron dot structures.

Use a table of electronegativity values to predict the following:  
the type of bond formed.

Which atom will attract the electron the most and take on a negative charge (polarity).

Why is water a polar molecule?

# English 3 Semester 2 Study Guide

**Michael Brown**



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